

# LIMITING REACTANT STOICHIOMETRY WORKSHEET

Name _____	Section _____	Date _____
<b>Limiting Reactants</b>		
<b>Directions:</b> Determine the limiting reactant in each problem.		
1) In $2\text{NaOH} + \text{MgCl}_2 \rightarrow \text{Mg(OH)}_2 + 2\text{NaCl}$ , 5.818 grams of NaOH react with 2.712 grams of $\text{MgCl}_2$ .		
2) 0.0718 grams of HCl react with 0.0002 grams of $\text{Ca(OH)}_2$ in the reaction $2\text{HCl} + \text{Ca(OH)}_2 \rightarrow 2\text{H}_2\text{O} + \text{CaCl}_2$ .		
3) 0.2000 grams of Nitrogen in the reaction $\text{N}_2 + 3\text{H}_2 \rightarrow 2\text{NH}_3$ combine with 0.2000 grams of hydrogen.		

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1) In $2\text{NaOH} + \text{MgCl}_2 \rightarrow \text{Mg(OH)}_2 + 2\text{NaCl}$ , 5.818 grams of NaOH react with 2.712 grams of $\text{MgCl}_2$ .		
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3) 0.2000 grams of Nitrogen in the reaction $\text{N}_2 + 3\text{H}_2 \rightarrow 2\text{NH}_3$ combine with 0.2000 grams of hydrogen.		
4) In the reaction $2\text{C}_2\text{H}_2 + 5\text{O}_2 \rightarrow 4\text{CO}_2 + 2\text{H}_2\text{O}$ , 10.000 grams of $\text{O}_2$ react with 5.908 grams of $\text{C}_2\text{H}_2$ .		
5) In the reaction $2\text{Al(OH)}_3 + 3\text{H}_2\text{CO}_3 \rightarrow \text{Al}_2(\text{CO}_3)_3 + 6\text{H}_2\text{O}$ , 1.18 grams of $\text{Al(OH)}_3$ and 1.312 grams of $\text{H}_2\text{CO}_3$ combine.		
6) 0.838 grams of sodium and 0.7021 grams of sodium nitrate react in the following process: $10\text{Na} + 2\text{NaNO}_3 \rightarrow 6\text{Na}_2\text{O} + \text{N}_2$ .		
7) 0.0700 grams of $\text{AsH}_3$ react with 1.201 grams of $\text{H}_2\text{SO}_4$ to produce products in the following reaction: $2\text{AsH}_3 + 3\text{H}_2\text{SO}_4 \rightarrow 2\text{AsH} + \text{As}_2\text{PO}_4$ .		
8) When $\text{K}_3\text{PO}_4$ and NaOH react in $\text{K}_3\text{PO}_4 + 3\text{NaOH} \rightarrow 3\text{KOH} + \text{Na}_3\text{PO}_4$ , 5.41 grams of NaOH react with 3.88 grams $\text{K}_3\text{PO}_4$ .		
9) In $\text{Pb(NO}_3)_2 + \text{H}_2\text{S} \rightarrow \text{PbS} + \text{HNO}_3$ , 10.02 grams of $\text{Pb(NO}_3)_2$ react with 7.98 grams of $\text{H}_2\text{S}$ .		
10) 0.4000 grams of $\text{H}_2\text{SO}_4$ combine with 0.2880 grams of $\text{SO}_2$ in the reaction $2\text{H}_2\text{SO}_4 + 2\text{SO}_2 \rightarrow 2\text{S}_2\text{O}_8^{2-} + 4\text{H}^+$ .		

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A CHEMISTRY WORKSHEET

10 PROBLEMS

WORKED ANSWER KEY

# Limited Reagent Lab Key

**Karin Nielsen-Saines**



## Limited Reagent Lab Key:

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*Chemical Reactor Design and Technology* Hugo de Lasa,2012-12-06 Today s frustrations and anxieties resulting from two energy crises in only one decade show us the problems and fragility of a world built on high energy consumption accustomed to the use of cheap non renewable energy and to the acceptance of eXisting imbalances between the resources and demands of countries Despite all these stressing factors our world is still hesitating about the urgency of undertaking new and decisive research that could stabilize our future Could this trend change in the near future In our view two different scenarios are possible A renewed energy tension could take place with an unpredictable timing mostly related to political and economic factors This could bring again scientists and technologists to a new state of shock and awaken our talents A second interesting and beneficial scenario could result from the positive influence of a new generation of researchers that with or without immediate crisis acting both in industry and academia will face the challenge of developing technologies and processes to pave the way to a less vulnerable society Because Chemical Reactor Design and Technology activities are at the heart of these required new technologies the timeliness of the NATO Advanced Study Institute at the University of Western Ontario London was very appropriate

**General Chemistry Key Concise Review Guide Sixth Edition** Darrell D. Ebbing,1999

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ChemCom American Chemical Society, 1993      *SourceBook Version 2.1*, 1998      *Addison-Wesley Chemistry*, 1990  
*Semimicro Experiments in Organic Chemistry* Bruce Nelson Campbell, Ernest Ray Kline, 1964      *Exploring Chemistry (Loose-Leaf)* Matthew Johll, 2012-05-04  
Matthew Johll's *Exploring Chemistry* covers the standard topics for the nonmajors course in the typical order but each chapter unfolds in the context of a single case study that helps students connect what they are learning to real life situations. For example, students work through the often difficult topics of molecular structure, gas laws, and organic chemistry by learning about the development of powerful new chemotherapy drugs, new technologies for screening airline passengers, and the creation of biodegradable biopolymers. It's the same case-driven approach that Johll uses in his acclaimed *Investigating Chemistry* now in its Third Edition, but *Exploring Chemistry* goes beyond the other book's specific focus on examples from forensic science to use real life stories from cooking, athletics, genetics, green chemistry, and more.

**Chemistry** William R. Robinson, Jerome D. Odom, Henry Fuller Holtzclaw, 1992      *Chemistry* Neil D. Jespersen, Alison Hyslop, 2021-12-03  
*Chemistry: The Molecular Nature of Matter* 8th Edition continues to focus on the intimate relationship that exists between structure at the atomic/molecular level and the observable macroscopic properties of matter. Key revisions in this edition focus on three areas: The deliberate inclusion of more updated real world examples that relate common real world student experiences to the science of chemistry. Simultaneously, examples and questions have been updated to align them with career concepts relevant to the environmental, engineering, biological, pharmaceutical, and medical sciences. Providing students with transferable skills with a focus on integrating metacognition and three-dimensional learning into the text. When students know what they know, they are better able to learn and incorporate the material. Providing a total solution through New WileyPLUS by fully integrating the enhanced eText with online assessment, answer-specific responses, and additional practice resources. The 8th edition continues to emphasize the importance of applying concepts to problem solving to achieve high-level learning and increase retention of chemistry knowledge. Problems are arranged in an intuitive, confidence-building order.

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