

# Limiting Reactant Practice

5 pages + 5-page Answer KEY + Periodic Table

## Chemistry

**LIMITING REACTANTS**

1. The reaction below produces ammonia for use in fertilizers.

$$N_2 + 3H_2 \rightarrow 2NH_3$$

If you start with 28 grams of  $N_2$  and 12 grams of  $H_2$ , determine the limiting reactant.

2. Calculate the grams of ammonia produced.

3. How many grams of the excess reactant is present in a left over?

4. Multiply the mass of the limiting reactant by the mole ratio of the excess reactant.

5. Subtract mass of excess  $\times$  mole of product.

1. Write a balanced chemical equation for the reaction below and include the physical states of the products. (include grams, mole ratios, and mole ratio of products)

$$2NaOH + H_2SO_4 \rightarrow Na_2SO_4 + 2H_2O$$

If you have 40 grams of one reactant and 80 grams of the other reactant, determine the limiting reactant.

2. How many grams of water produced are produced?



1. Magnesium chloride is used to create salts for de-icing roads. The reaction with very impure sodium is produced.

$$Mg + 2NaCl \rightarrow MgCl_2 + 2Na$$

If you start with 20 grams of magnesium and 2 moles of sodium chloride, determine the limiting reactant.

2. Calculate the grams of magnesium chloride produced.

3. How many grams of the excess reactant is present in a left over?

1. Magnesium chloride is used to create salts for de-icing roads. The reaction with very impure sodium is produced.

$$Mg + 2NaCl \rightarrow MgCl_2 + 2Na$$

If you have with 20 grams of magnesium and 80 grams of sodium, determine the limiting reactant.

2. How many grams of sodium chloride is produced?

1. The reaction below produces ammonia for use in fertilizers and a common household product. (include grams, mole ratios, and mole ratio of products)

$$N_2 + 3H_2 \rightarrow 2NH_3$$

If you have 40 grams of nitrogen and 80 grams of hydrogen, determine the limiting reactant.

2. Calculate the grams of ammonia produced.

3. How many grams of the excess reactant is present in a left over?

4. Multiply the mass of the limiting reactant by the mole ratio of the excess reactant.

5. Subtract mass of excess  $\times$  mole of product.

# ANSWER KEY INCLUDED

# Limiting Reactant Problems Answer Key

**ML Morrison**



**Limiting Reactant Problems Answer Key:**

## Reviewing **Limiting Reactant Problems Answer Key**: Unlocking the Spellbinding Force of Linguistics

In a fast-paced world fueled by information and interconnectivity, the spellbinding force of linguistics has acquired newfound prominence. Its capacity to evoke emotions, stimulate contemplation, and stimulate metamorphosis is really astonishing. Within the pages of "**Limiting Reactant Problems Answer Key**," an enthralling opus penned by a very acclaimed wordsmith, readers attempt an immersive expedition to unravel the intricate significance of language and its indelible imprint on our lives. Throughout this assessment, we shall delve into the book's central motifs, appraise its distinctive narrative style, and gauge its overarching influence on the minds of its readers.

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