

2nd Year Engineering (COP) 19 and 20 class, Semester Fourth, 2021

Chem

AC/2021/2021

Write in Short:

- 1) Define the term **concentration** in solid solution. **2M**
 (any) **→** **It is the ratio of the mass of solute to the mass of solution or the concentration of the solute.**
- 2) Write the unit and symbol for **density**. **2M**
 (any) **→** **Unit: kg m⁻³ or g cm⁻³ Symbol: ρ**
- 3) How would you determine the **mass fraction** and **mole fraction** of a component in a mixture? **2M**
 (any) **→** **Mass fraction = mass of component / total mass of mixture
 Mole fraction = moles of component / total moles of mixture**
- 4) How are **weight percentage** and **mole percentage** defined? **2M**
 (any) **→** **Weight percentage = (mass of component / total mass) × 100
 Mole percentage = (moles of component / total moles) × 100**
- 5) What are the **molar** and **mass** percentages? **2M (2M each)**
 (any) **→** **Molar percentage = (moles of component / total moles) × 100
 Mass percentage = (mass of component / total mass) × 100**
- 6) What is the **difference** between **weight percentage** and **mole percentage**? **2M**
 (any) **→** **Weight percentage is based on mass, while mole percentage is based on moles. Weight percentage is used for mixtures of solids, while mole percentage is used for mixtures of gases.**
- 7) What is the **basic** assumption in **ideal solution**? **2M**
 (any) **→** **The basic assumption in an ideal solution is that the interactions between the molecules of the solute and the solvent are similar to the interactions between the molecules of the pure components.**
- 8) How would you define the **partial molar volume** of a component in a mixture? **2M**
 (any) **→** **Partial molar volume is the volume occupied by one mole of a component in a mixture, at constant temperature and pressure.**
- 9) What is the **relationship** between **partial molar volume** and **molar volume**? **2M**
 (any) **→** **Partial molar volume is the molar volume of a component in a mixture, at constant temperature and pressure.**
- 10) How would you define the **partial molar volume** of a component in a mixture? **2M**
 (any) **→** **Partial molar volume is the volume occupied by one mole of a component in a mixture, at constant temperature and pressure.**
- 11) **Discuss** the **partial molar volume** of a component in a mixture. **2M**
 (any) **→** **Partial molar volume is the volume occupied by one mole of a component in a mixture, at constant temperature and pressure. It is a function of composition and temperature.**

Statement	Accounting of units in the equation	Final result in the equation
Spring constant k is defined as the force per unit extension.	$[k] = \frac{[F]}{[x]}$	$[k] = \frac{[MLT^{-2}]}{[L]} = [MT^{-2}]$
The force constant k is defined as the force per unit extension.	$[k] = \frac{[F]}{[x]}$	$[k] = \frac{[MLT^{-2}]}{[L]} = [MT^{-2}]$
Force constant k is defined as the force per unit extension.	$[k] = \frac{[F]}{[x]}$	$[k] = \frac{[MLT^{-2}]}{[L]} = [MT^{-2}]$

Icp Final Exam Review Answers

Thomas Griffiths



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