

CHAPTER 9

Acids and Bases

If the chemist has lost his soul, he will not have lost his courage and as he descends into the inferno and smells the homey fumes of brimstone, he will call out, "Amosdeus, hand me a test tube."

Gilbert Newton Lewis

We have already looked at several ways of classifying substances. Another way to classify substances is by their pH. Substances can be:

- Acid
- Base
- Neutral

9.1 Indicators

An indicator is a substance which is one colour in an acid and another colour in a base. It can therefore be used to test for the presence of an acid or a base. Some indicators are available in paper form, e.g. litmus and universal indicator. Others are available in a liquid form, e.g. methyl orange and phenolphthalein.

The colour changes of some common indicators are summarised in Table 9.1.

Table 9.1
Some common indicators

Indicator	Colour in Acid	Colour in Base
Litmus	Red	Blue
Phenolphthalein	Colourless	Pink
Methyl orange	Red	Yellow
Screened methyl orange	Red	Green
Bromothymol blue	Yellow	Blue

9.2 The pH Scale

The strength of an acid or base is indicated by its pH. The pH scale is a number line which runs from 0 to 14. Universal indicator can give a different colour for each pH.

Id A Chapter 9 Acids And Bases

Spencer Seager, Michael R Slabaugh



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Acid-Base, Fluids, and Electrolytes Made Ridiculously Simple Richard A. Preston, M.D., 2025-12-24 Useful for medical students interns and residents hospitalists ICU caretakers nurses responsible for IV fluid therapy physician associates and first year nephrology fellows A brief highly readable book providing the clinician with a straightforward approach to solving even the most complex acid base fluids and electrolyte problems Begins with the basic physiology that is key to understanding clinical water electrolyte and acid base disorders Numerous case examples Topics Include The Basics IV Solutions and IV Orders Hyponatremia Hyponatremia Hypokalemia Hyperkalemia Metabolic Acidosis Metabolic Alkalosis Three Step Diagnosis of Acid Base Disorders Case Examples [Acid-Base Balance and Nitrogen Excretion in Invertebrates](#)

Dirk Weihrauch, Michael O'Donnell, 2016-12-15 This textbook provides a comprehensive overview on the diverse strategies invertebrate animals have developed for nitrogen excretion and maintenance of acid base balance and summarizes the most recent findings in the field obtained by state of the art methodology A broad range of terrestrial freshwater and marine invertebrate groups are covered including crustaceans cephalopods insects and worms In addition the impact of current and future changes in ocean acidification on marine invertebrates due to anthropogenic CO₂ release will be analyzed The book addresses graduate students and young researchers interested in general animal physiology comparative physiology and marine aquatic animal physiology Also it is an essential source for researchers dealing with the effects of increasing pCO₂ levels on aquatic animals of which the vast majority are indeed invertebrates All chapters are peer reviewed

Chemical Principles Peter Atkins, Loretta Jones, 2007-08 Written for calculus inclusive general chemistry courses Chemical Principles helps students develop chemical insight by showing the connections between fundamental chemical ideas and their applications Unlike other texts it begins with a detailed picture of the atom then builds toward chemistry's frontier continually demonstrating how to solve problems think about nature and matter and visualize chemical concepts as working chemists do Flexibility in level is crucial and is largely established through clearly labeling separating in boxes the calculus coverage in the text Instructors have the option of whether to incorporate calculus in the coverage of topics The multimedia integration of Chemical Principles is more deeply established than any other text for this course Through the unique eBook the comprehensive Chemistry Portal Living Graph icons that connect the text to the Web and a complete set of animations students can take full advantage of the wealth of resources available to them to help them learn and gain a deeper understanding

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Chemical Equilibrium and Analysis Richard W. Ramette, 1981 Introduction to Acid-base and Electrolyte Balance Gösta Rooth, 1969

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