

Ideal Gas Law Practice Worksheet

Solve the following problems using the ideal gas law:

- 1) How many moles of gas does it take to occupy 120 liters at a pressure of 2.3 atmospheres and a temperature of 340 K?

- 2) If I have a 50 liter container that holds 45 moles of gas at a temperature of 200° C, what is the pressure inside the container?

- 3) It is not safe to put aerosol canisters in a campfire, because the pressure inside the canisters gets very high and they can explode. If I have a 1.0 liter canister that holds 2 moles of gas, and the campfire temperature is 1400° C, what is the pressure inside the canister?

- 4) How many moles of gas are in a 30 liter scuba canister if the temperature of the canister is 300 K and the pressure is 200 atmospheres?

- 5) I have a balloon that can hold 100 liters of air. If I blow up this balloon with 3 moles of oxygen gas at a pressure of 1 atmosphere, what is the temperature of the balloon?

Ideal Gas Law Problems Answer Key

Matthew Johll



Ideal Gas Law Problems Answer Key:

Math Concepts for Food Engineering Richard W. Hartel, D.B. Hyslop, Terry Howell, Jr., 1997-05-19 Food industry managers, engineers, and scientists require a personal repertory of mathematical tools to find solutions to a variety of work problems. These problems relate to processing, research, analysis, and other operations that require mathematical solutions. This new handbook provides a practical review of required basic mathematical methods and relates them to specific areas of food industry problem solving. Prepared for students in undergraduate food science programs, the handbook also serves the needs of food industry technical personnel who would be helped by a review of basic math for food engineering and non-technical managers who seek an introduction to food engineering math and problem solving. Each chapter on a math or problem solving skill begins with fundamental concepts, presents step-by-step procedures in the skill, and provides numerous practical illustrative examples. Practice problems with answers in the appendix provide for self-testing and reinforcement of learning. All of the math skills presented relate to common food industry problems. From the text: In food engineering and any other mathematically oriented discipline, one must become adept at problem solving using equations and mathematical principles. Many of the problems that food scientists and engineers face require the ability to apply physical and mathematical principles to solve equations and calculate necessary information. Thus, problem solving abilities are extremely important, and we should be familiar with the general steps in problem solving.

Investigating Chemistry Matthew Johll, 2008-12-22 In its new second edition, *Investigating Chemistry: A Forensic Science Perspective* remains the only book that uses the inherently fascinating topics of crime and criminal investigations as a context for teaching the fundamental chemical concepts most often covered in an introductory nonmajors course. Covering all the standard topics, Matthew Johll capitalizes on the surge of interest in the scientific investigation of crime as sparked by CSI and other television shows, bringing together the theme of forensic science and the fundamentals of chemistry in ways that are effective and accessible for students. This edition features refined explanations of the chemical concepts which are the core of the book, as well as a more thoroughly integrated forensic theme, updated features, and an expanded media supplements package.

Ebook: Chemistry: The Molecular Nature of Matter and Change Silberberg, 2015-01-16 *Ebook Chemistry: The Molecular Nature of Matter and Change*

Oil & Gas Journal, 1927

Chemistry: The Central Science Theodore L. Brown, H. Eugene LeMay Jr., Bruce E. Bursten, Catherine Murphy, Patrick Woodward, Steven Langford, Darius Sagatys, Adrian George, 2013-10-04 If you think you know the Brown LeMay Bursten Chemistry text, think again. In response to market request, we have created the third Australian edition of the US bestseller *Chemistry: The Central Science*. An extensive revision has taken this text to new heights. Triple-checked for scientific accuracy and consistency, this edition is a more seamless and cohesive product yet retains the clarity, innovative pedagogy, functional problem solving, and visuals of the previous version. All artwork and images are now consistent in quality across the entire text. And with a more traditional and logical organization of the Organic

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University Physics Kenneth E. Jesse, 1987 *How to Solve Physics Problems* Daniel Milton Oman, Robert Milton Oman, 2016-01-01

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