

Mixed Stoichiometry Problems

Enter only the number value for your answer. **NO UNITS**. If you enter units, it will not grade your responses correctly. **Round to the nearest whole number**. You must **show your t-chart** to earn credit for correct responses.



- How many moles of H_2 would be required to produce 5.0 moles of water?
- What mass of H_2O is formed when H_2 reacts with 384 g of O_2 ?
- How many moles are there in 67.2 grams of H_2 ?



- If 30.0 g of H_2SO_4 reacts with sodium hydroxide, what mass of water is produced?
- What mass of H_2SO_4 would be required to react with 0.75 mol of NaOH ?
- Determine the theoretical yield of water when 11.2 g of H_2SO_4 reacts with 7.1 g of NaOH .



- How many grams of SnF_2 are produced from the reaction of 30.0 g of HF ?
- How many moles of HF will be needed to react with 48 grams of Tin?
- How many moles are in 4.12×10^{24} atoms of HF ?

4. In a spacecraft, the carbon dioxide exhaled by astronauts can be removed by its reaction with lithium hydroxide, LiOH , according to the following chemical equation.



- If 20.0 moles of CO_2 is exhaled, the average amount exhaled by a person each day, how much (in grams) $\text{Li}_2\text{CO}_3(\text{s})$ is produced?

Mixed Stoichiometry Practice

Julie C. Gilbert

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