

Chemistry - Nail Lab

Purpose

The purpose is to determine the ratio of copper produced to iron consumed in a replacement reaction.

Procedure

Day 1

1. Label, then mass a 250 mL beaker.
2. Put between 3.0 and 4.0 g of copper (II) Chloride in the beaker.
3. Add about 50 mL distilled water to the beaker. Stir to dissolve the solid.
4. Mass 1 or 2 nails together to closest ± 0.1 g.
5. Place the nails in the copper chloride solution. Observe the reaction; record your observations. Place the labeled beaker in the place designated by your teacher.

Day 2

6. Remove the nails. Rinse or scrape the precipitate (copper metal) from the nails into your labeled 250 mL beaker. Place the nails in a labeled small beaker. Note the appearance of the nails.
7. Decant solution from the 250 mL beaker. Rinse the precipitate with about 25 mL of distilled water. Try to lose as little of the solid copper as you can when you decant. After the 2nd rinse with distilled water, rinse one last time with distilled water. Then place the labeled beaker for drying.
8. Mass the dry nails, then discard the nails.

Day 3

9. Mass the beaker + dry copper. Discard the copper in the place designated by your teacher. Wash your beaker and let dry.

Modeling Chemistry Nail Lab

Michael Seilmaier



Modeling Chemistry Nail Lab:

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