

## Unit 6 – Worksheet 3

### Ionic Compounds

#### Properties

#### Basic structural unit

1. Give the name of the following simple binary ionic compounds.

- $\text{Na}_2\text{O}$
- $\text{K}_2\text{S}$
- $\text{MgCl}_2$
- $\text{CaBr}_2$
- $\text{BaI}_2$
- $\text{Al}_2\text{S}_3$
- $\text{CaBr}$
- $\text{AgF}$

2. Give the name of the following simple binary ionic compounds.

- $\text{Na}_3\text{N}$
- $\text{K}_2\text{O}$
- $\text{AgBr}$
- $\text{MgI}_2$
- $\text{SrO}$

3. Write the formula for the following binary ionic compounds.

- lithium bromide
- sodium iodide
- silver sulfide
- cesium oxide
- beryllium iodide
- barium hydride
- aluminum fluoride
- potassium oxide

4. Write the formula for these ionic substances.

- silver oxide
- aluminum sulfide
- sodium nitride
- barium chloride
- strontium hydride

# Modeling Chemistry Unit 6 Test

**Michal Rosen-Zvi**



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range and information content of the toxicity endpoints detected and therefore the relevance of test systems and data interpretation while new techniques for non invasive diagnostic imaging and high resolution detection methods have permitted an increased role for human studies Several key examples of how these technologies are being harnessed to meet 21st century safety assessment challenges are provided including their deployment in integrated testing schemes in conjunction with kinetic modelling and in specialised areas such as inhalation toxicity studies

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