

# Chemistry: Final Exam Review With 100% Correct Answers 2024

How many valence electrons are in an atom of chlorine? - answer✓✓7

How many valence electrons does aluminum have? - answer✓✓3

Define Reaction Rate - answer✓✓The decrease in concentration of reactants or the increase in concentration of products with time.

How does increasing temperature affect the rate of reaction? - answer✓✓Increasing temperature causes the kinetic energy to increase. Thus, molecules are moving faster and colliding more frequently. This causes the rate of reaction to increase.

If I decrease the concentration of reactant, what happens to the rate of reaction? - answer✓✓Decreasing concentration leads to less reactants available for collision. Therefore, reaction rate decreases.

How does surface area influence reaction rate? - answer✓✓By increasing the surface area of reactant molecules, we are increasing the space or point of contact that these molecules have available to interact. This causes an increase in the rate of reaction.

How does a catalyst speed up a chemical reaction? - answer✓✓A catalyst speeds up a chemical reaction by finding an alternate pathway with a lower activation energy.

Define activation energy. - answer✓✓The energy required to start a chemical reaction.

Define equilibrium. - answer✓✓Equilibrium is established when the rate of the forward reaction is equal to the rate of the reverse reaction.

Is energy absorbed or released in an endothermic reaction? - answer✓✓Absorbed.

What is the sign of delta H for an exothermic reaction? - answer✓✓Exothermic reactions release energy. The sign of delta H is negative.

When you touch a flask and it feels hot, is this an endothermic or exothermic process? - answer✓✓Exothermic; heat is being given off.

Define specific heat capacity. - answer✓✓The amount of heat required to raise 1g of a substance to 1 degrees Celsius.

# Modeling Chemistry Final Exam Review Answers

**Shawn R. Simonson**



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to assist this process conduct research expanding what is known about learning and teaching and provide professional development and collegiality from elementary teachers to college professors As a pedagogy it has been shown to be effective in a variety of content areas and at different educational levels This is an introduction to the process and the community Every POGIL classroom is different and is a reflection of the uniqueness of the particular context the institution department physical space student body and instructor but follows a common structure in which students work cooperatively in self managed small groups of three or four The group work is focused on activities that are carefully designed and scaffolded to enable students to develop important concepts or to deepen and refine their understanding of those ideas or concepts for themselves based entirely on data provided in class not on prior reading of the textbook or other introduction to the topic The learning environment is structured to support the development of process skills such as teamwork effective communication information processing problem solving and critical thinking The instructor s role is to facilitate the development of student concepts and process skills not to simply deliver content to the students The first part of this book introduces the theoretical and philosophical foundations of POGIL pedagogy and summarizes the literature demonstrating its efficacy The second part of the book focusses on implementing POGIL covering the formation and effective management of student teams offering guidance on the selection and writing of POGIL activities as well as on facilitation teaching large classes and assessment The book concludes with examples of implementation in STEM and non STEM disciplines as well as guidance on how to get started Appendices provide additional resources and information about The POGIL Project

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