

5. In the copper–silver nitrate lab copper metal and silver nitrate solution reacted to produce silver metal and copper(II) nitrate in solution. A student placed a copper wire with a mass of 2.93 g in the reaction test tube. The silver nitrate solution contained 1.41 g of silver nitrate. He obtained 0.87 g of silver metal. Calculate the percent yield of silver. Also calculate the ending mass of all substances.

6. When hydrochloric acid (HCl) is added to sodium hydrogen carbonate, the products are water, aqueous sodium chloride and carbon dioxide gas. What is the per cent yield if 4.68 g of CO_2 are collected when 10.0 g of sodium hydrogen carbonate reacts with excess HCl? Also calculate the ending mass of all substances.

7. Phosphorus and bromine react vigorously together to form phosphorus tribromide. If 5.0 g of phosphorus and 35 g of bromine react, how many grams of PBr_3 could be produced? Also calculate the ending mass of all substances.

Modeling Chemistry U7 Ws 2 V2 0

JL Elias



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