

$$\text{percent yield} = \frac{\text{actual yield (g)}}{\text{theoretical yield (g)}} \times 100\%$$

↳ from B-C-A chart

Stoichiometry Worksheet 2: Percent Yield

For each of the problems below:

- Write the balanced chemical equation.
- Identify the given (with units) and what you want to find (with units).
- Show set up with units. Give final answer with units and label.

1. Using the Hoffman apparatus for electrolysis, a chemist decomposes 36 g of water into its gaseous elements. How many grams of hydrogen gas should she get (theoretical yield)?



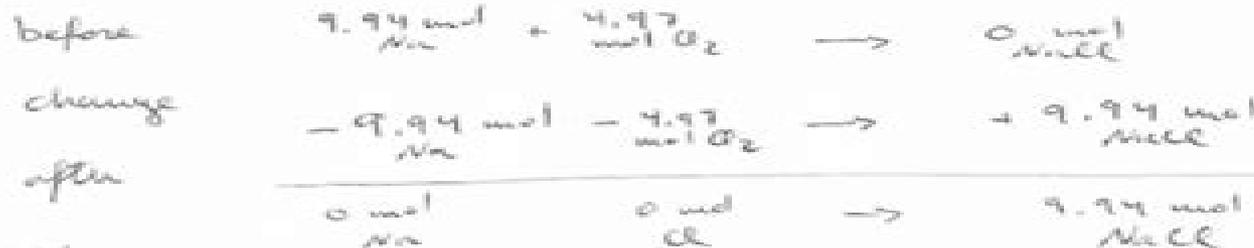
$$1 \text{ mol H}_2\text{O} : 18.02 \text{ g/mol}$$

$$36 \text{ g} \times \frac{1 \text{ mol H}_2\text{O}}{18.02 \text{ g}} = 1.99 \text{ mol H}_2\text{O}$$

$$1 \text{ mol H}_2 : 2.02 \text{ g}$$

$$2 \text{ mol H}_2 = \boxed{4.04 \text{ g H}_2}$$

2. Recall that liquid sodium reacts with chlorine gas to produce sodium chloride. You want to produce 581 g of sodium chloride. How many grams of sodium are needed?



$$1 \text{ mol NaCl} = 58.44 \text{ g}$$

$$581 \text{ g NaCl} \times \frac{1 \text{ mol NaCl}}{58.44 \text{ g}}$$

$$= \underline{9.94 \text{ mol NaCl}}$$

$$1 \text{ mol Na} = 22.99 \text{ g}$$

$$9.94 \text{ mol Na} \times 22.99 \text{ g} = \boxed{228.52 \text{ g Na}}$$

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Marcel A. Müller



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