

## Chemistry Unit 8 Worksheet 4

### Samples of Every Kind of Problem

*On a separate sheet of paper, write a complete solution to each of the problems below. Follow the procedure outlined in class. Be sure to circle your final answer.*

1. Calculate the number of moles of potassium chlorate,  $\text{KClO}_3$  (s), that must decompose to produce potassium chloride,  $\text{KCl}$  (s), and 1.8 moles of oxygen gas.

<b>Equation:</b>	$2 \text{KClO}_3 \rightarrow 2 \text{KCl} + 3 \text{O}_2$		
<b>Before:</b>	1.2	0	0
<b>Change:</b>	-1.2	+1.2	+1.8
<b>After:</b>	0	1.2	1.8

1.2 moles  
 $\text{KClO}_3$  needed

2. In a single displacement reaction, magnesium metal reacts with hydrochloric acid to produce magnesium chloride and hydrogen gas. How many moles of hydrochloric acid are needed to completely react with 2.43 g of magnesium?

<b>Equation:</b>	$\text{Mg} + 2 \text{HCl} \rightarrow \text{MgCl}_2 + \text{H}_2$		
<b>Before:</b>	0.100	?	0
<b>Change:</b>	-0.100	-0.200 +	0.100 + 0.100
<b>After:</b>	0	?	0.100 0.100

0.200 moles  
 HCl needed

3. Ethane,  $\text{C}_2\text{H}_6$  reacts with oxygen gas to produce carbon dioxide gas and water vapor. What mass of oxygen gas is required to react with 2.20 moles of ethane?

<b>Equation:</b>	$2 \text{C}_2\text{H}_6 + 7 \text{O}_2 \rightarrow 4 \text{CO}_2 + 6 \text{H}_2\text{O}$		
<b>Before:</b>	2.20	xs	0
<b>Change:</b>	-2.20	-7.70	+4.40 +6.60
<b>After:</b>	0	xs	4.40 6.60

$$\text{O}_2 \text{ needed: } 7.70 \text{ mole} \times \frac{32.0 \text{ g}}{1 \text{ mole}} = 246 \text{ g O}_2 \text{ required}$$

4. Determine the mass of sodium nitrate produced when 0.73 g of nickel (II) nitrate reacts with sodium hydroxide according to the following equation:



$$0.73 \text{ g} \times \frac{1 \text{ mole}}{183 \text{ g}} = .0040 \text{ mole}$$

<b>Equation:</b>	$\text{Ni}(\text{NO}_3)_2 + 2 \text{NaOH} \rightarrow \text{Ni}(\text{OH})_2 + 2 \text{NaNO}_3$		
<b>Before:</b>	0.0040	XS	0
<b>Change:</b>	-0.0040	-0.0080	+0.0040 +0.0080
<b>After:</b>	0	XS	0.0040 0.0080

$$\text{NaNO}_3 \text{ produced: } 0.0080 \text{ mole} \times \frac{85.0 \text{ g}}{1 \text{ mole}} = 0.68 \text{ g}$$

# Modeling Chemistry Unit 8 4

**G Orfield**



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