

8.1 Covalent bonding and covalent compounds

1. A **covalent bond** forms by the sharing of outermost shell electrons between two atoms. For example,

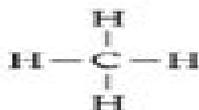


2. Covalent bond refers to the electrostatic attraction between the shared electrons (negatively charged) and the two nuclei (positively charged) of the bonded atoms.
3. A shared pair of electrons (a **bond pair of electrons** or a **bond pair**) makes a single covalent bond, e.g. H–H.
Two shared pairs of electrons make a double covalent bonds, e.g. O=O.
Three shared pairs of electrons make a triple covalent bond, e.g. N≡N.
4. Some atoms have unshared pairs of outermost shell electrons, which are known as **lone pairs of electrons** (or **lone pairs**), e.g.



8.2 Molecular formulae, structural formulae and molecular models

5. The **molecular formula** of a molecular substance shows the actual number of each kind of atoms in one molecule of the substance, e.g. CH₄.
6. The **structural formula** of a molecular substance shows the way that the atoms are joined together in one molecule of the substance, e.g.



7. **Molecular models**, including **ball-and-stick models** and **space-filling models**, can be used to show how the bonded atoms in molecules are arranged in space.

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