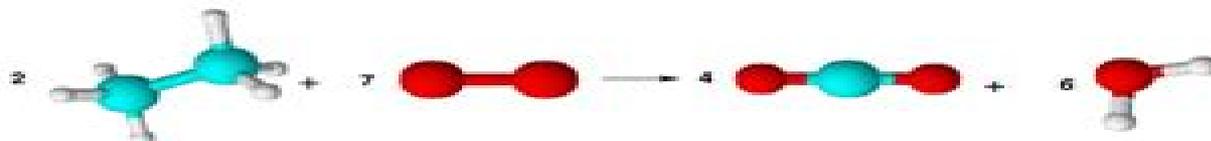


Practice Exercises

- 1.1 (a) SF_6 contains 1 S and 6 F atoms per molecule
 (b) $(\text{C}_2\text{H}_5)_2\text{N}_2\text{H}_2$ contains 4 C, 12 H, and 2 N per molecule
 (c) $\text{Ca}_3(\text{PO}_4)_2$ contains 3 Ca, 2 P, and 8 O atoms per formula unit
 (d) $\text{Co}(\text{NO}_3)_2 \cdot 6\text{H}_2\text{O}$ contains 1 Co, 2 N, 12 O, and 12 H per formula unit
- 1.2 (a) NH_4NO_3 contains 2 N nitrogen, 4 H hydrogen, 3 O oxygen atoms per formula unit
 (b) $\text{FeNH}_4(\text{SO}_4)_2$ contains 1 Fe iron, 1 N nitrogen, 4 H hydrogen, 2 S sulfur, 8 O oxygen atoms per formula unit
 (c) $\text{Mo}(\text{NO}_3)_2 \cdot 5\text{H}_2\text{O}$ contains 1 Mo molybdenum, 2 N nitrogen, 11 O oxygen, and 10 H hydrogen atoms per formula unit
 (d) $\text{C}_6\text{H}_4\text{ClNO}_2$ contains 6 C carbon, 4 H hydrogen, 1 Cl chlorine, 1 N nitrogen, and 2 O oxygen atoms per molecule
- 1.3 $\text{C}_2\text{H}_5\text{N}$, CH_3NHCH_3
- 1.4 Reactants: 4 N, 12 H, and 6 O; Products: 4 N, 12 H, and 6 O.
- 1.5 Reactants: 6 N, 42 H, 2 P, 20 O, 3 Ba, 12 C; Products: 3 Ba, 2 P, 20 O, 6 N, 42 H, 12 C; The reaction is balanced.

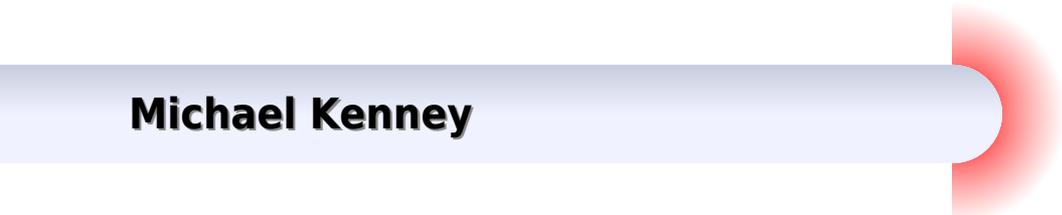


Review Questions

- 1.1 This answer will be student dependent.
- 1.2 Observation, testing and explanation.
- 1.3 (a) A law is a description of behavior based on the results of many experiments which are true while a theory is a tested explanation of the results of many experiments.
 (b) An observation is a statement that accurately describes something we see, hear, taste, feel or smell while a conclusion is a statement that is based on a series of observations.
 (c) Data are the observations made while performing experiments.
- 1.4 A theory is valid as long as there is no experimental evidence to disprove it. Any experimental evidence that contradicts the theory therefore, disproves the theory.
- 1.5 Matter has mass and occupies space. All items in the question are examples of matter.
- 1.6 A physical change does not change the chemical composition of matter. Melting, boiling, change of shape, or mass, and the formation of a mixture are examples of physical changes to matter.
 A chemical change changes the chemical composition of matter. Formation of new compounds from the reaction of other substances is an example.

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