

Mixed Gas Laws Worksheet - Solutions

- 1) How many moles of gas occupy 98 L at a pressure of 2.8 atmospheres and a temperature of 292 K?

$$n = \frac{PV}{RT} = \frac{(2.8 \text{ atm})(98 \text{ L})}{(0.0821 \text{ L}\cdot\text{atm}/\text{mol}\cdot\text{K})(292 \text{ K})} = 11 \text{ moles of gas}$$

- 2) If 5.0 moles of O_2 and 3.0 moles of N_2 are placed in a 30.0 L tank at a temperature of 25°C, what will the pressure of the resulting mixture of gases be? 25°C = 298 K

$$\text{O}_2: P = \frac{nRT}{V} = \frac{(5.0 \text{ mol})(0.0821 \text{ L}\cdot\text{atm}/\text{mol}\cdot\text{K})(298 \text{ K})}{(30.0 \text{ L})} = 4.1 \text{ atm}$$

$$\text{N}_2: P = \frac{nRT}{V} = \frac{(3.0 \text{ mol})(0.0821 \text{ L}\cdot\text{atm}/\text{mol}\cdot\text{K})(298 \text{ K})}{(30.0 \text{ L})} = 2.4 \text{ atm}$$

$$P_{\text{Total}} = P_{\text{O}_2} + P_{\text{N}_2} = 4.1 \text{ atm} + 2.4 \text{ atm} = 6.5 \text{ atm}$$

- 3) A balloon is filled with 35.0 L of helium in the morning when the temperature is 20.0°C. By noon the temperature has risen to 45.0°C. What is the new volume of the balloon?

$$T_1 = 20.0^\circ \text{C} = 293 \text{ K}, V_1 = 35.0 \text{ L}, T_2 = 45.0^\circ \text{C} = 318 \text{ K}, V_2 = ?$$

$$V_2 = \frac{V_1 T_2}{T_1} = \frac{(35.0 \text{ L})(318 \text{ K})}{(293 \text{ K})} = 38.0 \text{ L}$$

- 4) A 35 L tank of oxygen is at 315 K with an internal pressure of 190 atmospheres. How many moles of gas does the tank contain?

$$n = \frac{PV}{RT} = \frac{(190 \text{ atm})(35 \text{ L})}{(0.0821 \text{ L}\cdot\text{atm}/\text{mol}\cdot\text{K})(315 \text{ K})} = 260 \text{ moles of gas}$$

- 5) A balloon that can hold 85 L of air is inflated with 3.5 moles of gas at a pressure of 1.0 atmosphere. What is the temperature in °C of the balloon?

$$T = \frac{PV}{nR} = \frac{(1 \text{ atm})(85 \text{ L})}{(3.5 \text{ mol})(0.0821 \text{ L}\cdot\text{atm}/\text{mol}\cdot\text{K})} = 296 \text{ K} = 23^\circ \text{C}$$

- 6) CaCO_3 decomposes at 1200°C to form CO_2 gas and CaO . If 25 L of CO_2 are collected at 1200°C, what will the volume of this gas be after it cools to 25°C?

$$T_1 = 1200^\circ \text{C} = 1473 \text{ K}, V_1 = 25 \text{ L}, T_2 = 25^\circ \text{C} = 298 \text{ K}, V_2 = ?$$

$$V_2 = \frac{V_1 T_2}{T_1} = \frac{(25 \text{ L})(298 \text{ K})}{(1473 \text{ K})} = 5.1 \text{ L}$$

- 7) A helium balloon with an internal pressure of 1.00 atm and a volume of 4.50 L at 20.0°C is released. What volume will the balloon occupy at an altitude where the pressure is 0.600 atm and the temperature is -20.0°C?

$$P_1 = 1.00 \text{ atm}, V_1 = 4.50 \text{ L}, T_1 = 20.0^\circ \text{C} = 293 \text{ K}, P_2 = 0.600 \text{ atm}, V_2 = ?, T_2 = -20.0^\circ \text{C} = 253 \text{ K}$$

$$V_2 = \frac{P_1 V_1 T_2}{T_1 P_2} = \frac{(1.00 \text{ atm})(4.50 \text{ L})(253 \text{ K})}{(293 \text{ K})(0.600 \text{ atm})} = 6.48 \text{ L}$$

- 8) There are 135 L of gas in a container at a temperature of 260°C. If the gas was cooled until the volume decreased to 75 L, what would the temperature of the gas be?

$$T_1 = 260^\circ \text{C} = 533 \text{ K}, V_1 = 135 \text{ L}, T_2 = ?, V_2 = 75 \text{ L}$$

$$T_2 = \frac{V_2 T_1}{V_1} = \frac{(75 \text{ L})(533 \text{ K})}{(135 \text{ L})} = 296 \text{ K} = 23^\circ \text{C}$$

Mixed Gas Laws Key

Liyong Dong



Mixed Gas Laws Key:

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In a world taken by displays and the ceaseless chatter of instant communication, the melodic beauty and psychological symphony produced by the written term often diminish into the background, eclipsed by the persistent noise and disruptions that permeate our lives. Nevertheless, set within the pages of **Mixed Gas Laws Key** a marvelous literary value full of natural thoughts, lies an immersive symphony waiting to be embraced. Crafted by a masterful composer of language, this captivating masterpiece conducts viewers on a mental journey, well unraveling the concealed songs and profound influence resonating within each carefully crafted phrase. Within the depths of the moving analysis, we can explore the book's main harmonies, analyze its enthralling writing design, and submit ourselves to the profound resonance that echoes in the depths of readers' souls.

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