

Mixed Mole Problems Worksheet

Solve the following problems. Be sure to show all your work!

1. Find the number of moles of silver ions
- 2.50 g
- of
- Ag^+
- ions in a solution of silver nitrate.

$$2.50 \text{ g Ag}^+ \times \frac{1 \text{ mol Ag}^+}{107.87 \text{ g Ag}^+} = 0.0232 \text{ mol Ag}^+ \quad \boxed{0.0232 \text{ mol}}$$

2. How many moles of
- CaCl_2
- are there in
- 2.50 kg
- of
- CaCl_2
- ?

$$2.50 \text{ kg CaCl}_2 \times \frac{1000 \text{ g}}{1 \text{ kg}} \times \frac{1 \text{ mol CaCl}_2}{110.98 \text{ g CaCl}_2} = 22.5 \text{ mol CaCl}_2 \quad \boxed{22.5 \text{ mol CaCl}_2}$$

3. What is the mass of
- 0.025 mol
- of
- CaCl_2
- ?

$$0.025 \text{ mol CaCl}_2 \times \frac{110.98 \text{ g CaCl}_2}{1 \text{ mol CaCl}_2} = 2.77 \text{ g CaCl}_2 \quad \boxed{2.77 \text{ g CaCl}_2}$$

4. Find the number of moles of ions in
- 2.50 g
- of
- CaCl_2
- (molecular weight =
- 110.98 g/mol
-).

$$2.50 \text{ g CaCl}_2 \times \frac{1 \text{ mol CaCl}_2}{110.98 \text{ g CaCl}_2} \times \frac{3 \text{ mol ions}}{1 \text{ mol CaCl}_2} = 0.0675 \text{ mol ions} \quad \boxed{0.0675 \text{ mol ions}}$$

5. If you start off with
- 0.025 mol
- of
- CaCl_2
- and you add
- 0.025 mol
- of
- NaCl
- to that mixture, what is the total mass of this solution? Please show all your work.

$$\frac{0.025 \text{ mol}}{1 \text{ mol}} \times \frac{110.98 \text{ g}}{1 \text{ mol}} + \frac{0.025 \text{ mol}}{1 \text{ mol}} \times \frac{58.44 \text{ g}}{1 \text{ mol}} = 4.52 \text{ g} + 1.46 \text{ g} = 5.98 \text{ g}$$

6. What volume of
- 0.25 M
- CaCl_2
- solution contains
- 0.025 mol
- of
- CaCl_2
- ?

$$0.025 \text{ mol CaCl}_2 \times \frac{1 \text{ L solution}}{0.25 \text{ mol CaCl}_2} = 0.10 \text{ L solution} \quad \boxed{0.10 \text{ L solution}}$$

7. What is the mass of
- 0.025 mol
- of
- CaCl_2
- and
- 0.025 mol
- of
- NaCl
- ?

$$0.025 \text{ mol CaCl}_2 \times \frac{110.98 \text{ g}}{1 \text{ mol}} + 0.025 \text{ mol NaCl} \times \frac{58.44 \text{ g}}{1 \text{ mol}} = 2.77 \text{ g} + 1.46 \text{ g} = 4.23 \text{ g}$$

8. Calculate the molar mass of a gas of
- 0.25 M
- CaCl_2
- solution if
- 0.025 mol
- of
- CaCl_2
- is in
- 0.10 L
- of solution.

$$\frac{0.025 \text{ mol CaCl}_2}{0.10 \text{ L solution}} = \frac{0.025 \text{ mol}}{0.10 \text{ L}} = 0.25 \text{ mol/L} \quad \boxed{0.25 \text{ mol/L}}$$

Mixed Mole Problems Worksheet Key

LP Steffe



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