

Chemistry – Unit 10 Worksheet 3

1. Element X has two natural isotopes. The isotope with a mass number of 6 has a relative abundance of 7.5%. The isotope with a mass number of 7 has a relative abundance of 92.5%. Determine the average molar mass for the element from these figures. What is the true identity and atomic number of element X?
2. The element copper is found to contain the naturally occurring isotopes ${}^{63}_{29}\text{Cu}$ and ${}^{65}_{29}\text{Cu}$. The relative abundances are 69.1% and 30.9% respectively. Calculate the average molar mass of copper.
3. Uranium has three isotopes with the following percent abundances: ${}^{234}_{92}\text{U}$ (0.0058%), ${}^{235}_{92}\text{U}$ (0.71%), ${}^{238}_{92}\text{U}$ (99.23%). What do you expect the molar mass of uranium to be in whole numbers? Why?
4. A sample of silver as it occurs in nature is 52.0% of isotope ${}^{107}_{47}\text{Ag}$ and 48.0% of isotope ${}^{109}_{47}\text{Ag}$. What is the average molar mass of silver? (Compare your result with the value given in the periodic table).

Modeling Chemistry Unit 10 2 Answer

Ying Liu



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