

Unit 7: Balancing Chemical Reactions- Worksheet 2

Balance the following equations by inserting the proper coefficients.

- $2 \text{SO}_2 + 1 \text{O}_2 \rightarrow 2 \text{SO}_3$ **Synthesis**
 - $2 \text{CH}_4 + 3 \text{O}_2 \rightarrow 2 \text{CO} + 4 \text{H}_2\text{O}$ **Single R**
 - $2 \text{P} + 3 \text{Cl}_2 \rightarrow 2 \text{PCl}_3$ **Synthesis**
 - $2 \text{CO} + 1 \text{O}_2 \rightarrow 2 \text{CO}_2$ **Synthesis**
 - $2 \text{CH}_4 + 1 \text{O}_2 \rightarrow 2 \text{CH}_3\text{OH}$ **Synthesis**
 - $2 \text{Li} + 1 \text{Br}_2 \rightarrow 2 \text{LiBr}$ **Synthesis**
 - $2 \text{Al}_2\text{O}_3 \rightarrow 4 \text{Al} + 3 \text{O}_2$ **Decomp**
 - $2 \text{Na} + 2 \text{H}_2\text{O} \rightarrow 2 \text{NaOH} + 1 \text{H}_2$ **Single R**
 - $6 \text{CO}_2 + 6 \text{H}_2\text{O} \rightarrow 1 \text{C}_6\text{H}_{12}\text{O}_6 + 6 \text{O}_2$ **Rev. Comb**
 - $1 \text{H}_2\text{SO}_4 + 2 \text{NaCl} \rightarrow 2 \text{HCl} + 1 \text{Na}_2\text{SO}_4$ **Double R**
- Two reactions used to get rid of sulfur dioxide, a pollutant from burning coal:
- $3 \text{H}_2 + 1 \text{SO}_2 \rightarrow 1 \text{H}_2\text{S} + 2 \text{H}_2\text{O}$ **Double R**
 - $2 \text{CaCO}_3 + 2 \text{SO}_2 + 1 \text{O}_2 \rightarrow 2 \text{CaSO}_4 + 2 \text{CO}_2$ **No**
 - $2 \text{AgNO}_3 + 1 \text{CaCl}_2 \rightarrow 2 \text{AgCl} + 1 \text{Ca}(\text{NO}_3)_2$ **Double R**
 - $2 \text{HCl} + 1 \text{Ba}(\text{OH})_2 \rightarrow 1 \text{BaCl}_2 + 2 \text{H}_2\text{O}$ **Double R**
 - $1 \text{H}_3\text{PO}_4 + 3 \text{NaOH} \rightarrow 1 \text{Na}_3\text{PO}_4 + 3 \text{H}_2\text{O}$ **Double R**
 - $1 \text{Pb}(\text{NO}_3)_2 + 2 \text{KI} \rightarrow 1 \text{PbI}_2 + 2 \text{KNO}_3$ **Double R**
 - $3 \text{CuO} + 2 \text{NH}_3 \rightarrow 1 \text{N}_2 + 3 \text{Cu} + 3 \text{H}_2\text{O}$ **Single R**
 - $1 \text{C}_2\text{H}_5\text{OH} + 3 \text{O}_2 \rightarrow 2 \text{CO}_2 + 3 \text{H}_2\text{O}$ **Com**
 - $2 \text{C}_2\text{H}_6 + 3 \text{O}_2 \rightarrow 2 \text{CH}_3\text{COOH} + 2 \text{H}_2\text{O}$ **No**
 - $3 \text{NO}_2 + 1 \text{H}_2\text{O} \rightarrow 2 \text{HNO}_3 + 1 \text{NO}$ **No**

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