

# Empirical/Molecular Formula

Name: \_\_\_\_\_

Date: \_\_\_\_\_

1. What is an empirical formula? How is it different from a molecular formula?

**The empirical formula is the lowest whole number ratio of atoms in a compound. The molecular formula shows the true number of atoms in the molecule and is a whole number multiple of the empirical formula.**

2. State the empirical formula for each of the following.



3. A sample of a compound is 50.0% sulfur and 50.0% oxygen by mass. What is the empirical formula?

$$S: \frac{50.0 \text{ g S}}{32.07 \text{ g S}} \times \frac{1 \text{ mol S}}{1 \text{ mol S}} = \frac{1.559 \text{ mol S}}{1.559 \text{ mol}} = 1 S$$

**EF:  $SO_2$**

$$O: \frac{50.0 \text{ g O}}{16.00 \text{ g O}} \times \frac{1 \text{ mol O}}{1 \text{ mol O}} = \frac{3.125 \text{ mol O}}{1.559 \text{ mol}} = 2 O$$

4. A sample of a compound is 52.9% aluminum. The rest is oxygen. What is the empirical formula?

$$Al: \frac{52.9 \text{ g Al}}{26.98 \text{ g Al}} \times \frac{1 \text{ mol Al}}{1 \text{ mol Al}} = \frac{1.961 \text{ mol Al}}{1.961 \text{ mol Al}} = 1 Al \times 2 = 2$$

**EF:  $Al_2O_3$**

$$O: \frac{47.1 \text{ g O}}{16.00 \text{ g O}} \times \frac{1 \text{ mol O}}{1 \text{ mol O}} = \frac{2.944 \text{ mol O}}{1.961 \text{ mol}} = 1.5 \times 2 = 3 O$$

Ratio must be a whole

5. A compound is 48.65% carbon, 8.11 % hydrogen and 43.24 % oxygen. Determine the empirical formula.

4.05 mol carbon, 8.11 mol H, 2.70 mol O



1. Divide by the smallest # to get =  $C_{1.5}H_3O_1$

2. Multiply to get a whole number ratio (x2) =  $C_3H_6O_2$

6. Vinegar is a dilute solution of acetic acid. The molar mass of acetic acid is 60.0 g/mol and it has an empirical formula of  $CH_2O$ . What is the molecular formula of acetic acid?



7. A compound has an empirical formula of  $C_3H_4$ . Which of the following are possible molar masses of the compound? 20 g/mol, 55 g/mol, 80 g/mol, 120 g/mol.



# Empirical Formulas Answers

**Arun Bahl & B S Bahl**



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**CBSE Chemistry Chapterwise Case Study Class 11** Priti Singhal, 2024-11-17 This book is structured to align with the latest syllabus and curriculum guidelines ensuring that the content is both relevant and rigorous Each chapter begins with a clear set of learning objectives providing a roadmap for students to understand what they will achieve by the end of the chapter We have included numerous diagrams illustrations and real life examples to make complex concepts more accessible and engaging

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