

Chapter 6

End-of-Chapter Questions

1. A compound which, when dissolved in water, does not cause a current to flow, is called this.
2. Why is the simplest unit of an ionic compound called a formula unit and not a molecule?
3. When main-group metals lose all their outer-level electrons, what electron configuration do they obtain?
4. What holds the ions together in a crystal lattice?
5. When an atom of each of the following elements forms an ion, what will be its charge?
(a) potassium (b) sulfur (c) bromine
6. When an atom of each of the following elements forms an ion, what will be its charge?
(a) barium (b) phosphorus (c) aluminum
7. Identify the ion with the following combinations of protons (p) and electrons (e):
(a) 3 p, 2 e (b) 20 p, 18 e (c) 17 p, 18 e
8. Identify the ion with the following combinations of protons (p) and electrons (e):
(a) 26 p, 23 e (b) 14 p, 18 e (c) 30 p, 28 e
9. How many protons (p) and electrons (e) does each of the following species possess?
(a) Co^{2+} (b) Ag^+ (c) N^{3-}
10. How many protons (p) and electrons (e) does each of the following species possess?
(a) O^{2-} (b) Br^- (c) Ni^{2+}
11. Two species that have the same electronic structure are said to be isoelectronic. Identify the ion that is:
(a) isoelectronic with neon and has a charge of -4 ;
(b) isoelectronic with neon and a charge of $+3$.
12. Draw electron-dot symbols for the sodium ion and the iodide ion. Then deduce the ion-ratio of the ionic compound formed between the two ions and write its chemical formula.

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