

Lab 5: RATE AND ORDER OF A CHEMICAL REACTION: IODINATION OF ACETONE

I. Introduction

The purpose of this experiment is to study a reaction that, in the vicinity of room temperature, proceeds at a moderate, relatively easily measured rate. For a given reaction, the rate typically increases with an increase in the concentration of any reactant. The relation between rate and concentration is a remarkably simple one in many cases. The objectives for this experiment include to determine the orders of reaction with respect to acetone and hydrogen ion; to confirm that the order of reaction with respect to iodine is zero; and to evaluate the rate constant k for the mixtures prepared at room temperature. For this experiment, we took safety precautions using laboratory materials and solutions, including wearing personal protective equipment.

II. Results Analysis

A. Data

Table 1. Reaction Rate Determination

Mixture Number	Initial Concentrations*			Time for reaction (s)	Temperature (°C)**	Reaction rate (M/s or mole/L · s)
	Acetone (M)	Iodine (M)	HCl (M)			
I (D.1)	5 mL	5 mL	5 mL	71.92	25	
II (D.2)	10 mL	5 mL	5 mL	35.10	25	
III (D.3)	5 mL	5 mL	10 mL	39.11	25	
IV (D.4)	5 mL	2.5 mL	5 mL	42.54	25	

Data Table 1. Group C - Iodination of Acetone: Time to Complete Reaction

Mixture	Temperature (°C)	Time (s) Took for Mixture Color to Disappear (seconds)
Mixture 1 (D.1)	25	71.92
Mixture 2 (D.2)	25	35.10
Mixture 3 (D.3)	25	39.11
Mixture 4 (D.4)	25	42.54

B. Discussion

Based on my results, I observed how the freezing point depression is directly related to the molecular weight of the solute in the solution. Using this relationship, the molecular weight of

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M Woodhall



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