

# IONIC BONDS

**Ionic Bonds**

**Reading Passage**

Ionic bonds are a type of chemical bond formed due to the electrostatic attraction between oppositely charged ions. These ions are created when atoms transfer electrons, resulting in a cation (a positive ion) and a negative ion (anion). For instance, in sodium chloride (NaCl), sodium loses an electron to become a cation, and chlorine gains an electron to become an anion. The transfer of electrons creates ions that are attracted to each other because of their opposite charges.

The arrangement of ionic ions in compounds is often depicted as a crystal lattice—a highly organized, repeating three-dimensional pattern. This structure maximizes the electrostatic attractions between oppositely charged ions, which leads to characteristic physical properties such as high melting points, brittleness, and good solubility in water. The regular structure ensures that each ion is surrounded by ions of opposite charge, creating a stable configuration.

In the macroscopic world, the electrostatic attraction between ions results in a stable and highly ordered structure. This stability at the atomic level explains the distinctive properties of ionic compounds on a macroscopic scale. For example, the high melting points of ionic compounds are due to the significant energy required to overcome the strong electrostatic forces holding ions in the lattice. Additionally, ionic compounds are composed of a continuous network of oppositely charged ions, which makes them good conductors of electricity when dissolved in water or melted, as the ions are free to move and carry the current.

**True or False? Ionic Bond Concepts**

1. Ionic bonds form due to the electrostatic attraction between ions with opposite charges.	True
2. Sodium chloride sodium gains an electron to become a negative ion.	False
3. Ionic compounds are usually crystalline and have high melting points.	True
4. The organization of ions in compounds maximizes attraction between similar charges.	False
5. Electrostatic forces are the driving force behind ionic bond formation.	True
6. In the lattice structure, each ion is surrounded by ions of opposite charge.	True

# Ionic Bonds Guided Reading Answer Key

**Ray Boyington, William L. Masterton**

## **Ionic Bonds Guided Reading Answer Key:**

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