

Name: \_\_\_\_\_

Period: \_\_\_\_\_

## Fall Semester Review – IPC

IPC Chemistry Review

Find the atomic mass of the following compounds.

Calcium oxide (CaO) \_\_\_\_\_

Water \_\_\_\_\_

Classify (ionic, covalent or polyatomic) and name the following compounds.

1. CO<sub>2</sub> \_\_\_\_\_2. Na<sub>2</sub>O \_\_\_\_\_3. BeCrO<sub>4</sub> \_\_\_\_\_

Find the valence electrons and oxidation number of:

1. Helium \_\_\_\_\_ 2. Oxygen \_\_\_\_\_

Draw the Lewis Dot Diagrams for:

1. Carbon \_\_\_\_\_

2. Sodium \_\_\_\_\_

Make Balanced Compounds from:

1. Li + O \_\_\_\_\_

2. Na + Ar \_\_\_\_\_

3. Be + NO<sub>3</sub><sup>-</sup> \_\_\_\_\_

4. Ca + O \_\_\_\_\_

With dot diagrams draw the covalent bond of O<sub>2</sub>.With dot diagrams the covalent bond of O<sub>2</sub>.

Matching:

- A. Mixture
- B. Compound
- C. Element
- D. Matter
- E. Heterogeneous
- F. Homogeneous

- \_\_\_\_\_ A substance made up of two or more atoms that must be separated by chemical means.
- \_\_\_\_\_ Something made up of 2 or more substances that can be separated by physical means.
- \_\_\_\_\_ A substance that is pure and made up of only one type of atom.
- \_\_\_\_\_ Anything that has mass and takes up space.
- \_\_\_\_\_ A mixture that is the same throughout.
- \_\_\_\_\_ A mixture that is different throughout.

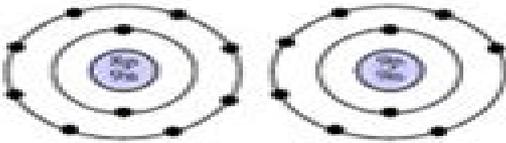
1. Valence electrons are the outermost electrons of an atom that are involved in chemical bonding? True/False.

2. Lithium has \_\_\_\_\_ valence electrons, will (lose or gain) \_\_\_\_\_ electrons and become (positive or negative). This is why lithium's oxidation number is \_\_\_\_\_.

3. A metal and non-metal will form a \_\_\_\_\_ compound, two non-metals form a \_\_\_\_\_ compound.

4. Protons are \_\_\_\_\_; electrons are \_\_\_\_\_; neutrons are \_\_\_\_\_.

5. Protons and neutrons are in the center of the atom, which is called the \_\_\_\_\_.



Which one is the ion?

Give the ion notation for the ion:

Atom A

Atom B

Give the number of protons for:

1. Carbon \_\_\_\_\_ 3. Iron \_\_\_\_\_

2. Beryllium \_\_\_\_\_ 4. Chlorine \_\_\_\_\_

Matching:

- A. Ion
- B. Element
- C. Isotope
- D. Oil

If you change the number of protons you change the \_\_\_\_\_.

If you change the number of neutrons, you change the \_\_\_\_\_.

If you change the number of electrons, you change the \_\_\_\_\_.

Every 5,000 miles you should change a car's \_\_\_\_\_.

Matching:

- A. Law of Conservation of Mass
- B. Meter
- C. Centimeter
- D. Liter
- E. Gram

\_\_\_\_\_ Mass is neither destroyed or created in chemical reactions.

\_\_\_\_\_ Unit of mass; about 1 dollar bill.

\_\_\_\_\_ 1/100th of meter; width of pinky finger.

\_\_\_\_\_ Unit of volume; just bigger a quart.

\_\_\_\_\_ Unit of length; 3.3.

Matching:

- A. Solid
- B. Liquid
- C. Gas
- D. Oil

\_\_\_\_\_ Molecules that are tightly packed and retain their shape and size.

\_\_\_\_\_ Molecules that bounce off of each other, can be compressed, and take the shape of their container.

\_\_\_\_\_ Molecules that can move (slide) over each other and have a definite size (volume), but not shape and can not be compressed.

# IPC Chemistry Review

**AN Whitehead**



## **IPC Chemistry Review:**

The book delves into Ipc Chemistry Review. Ipc Chemistry Review is an essential topic that needs to be grasped by everyone, ranging from students and scholars to the general public. This book will furnish comprehensive and in-depth insights into Ipc Chemistry Review, encompassing both the fundamentals and more intricate discussions.

1. The book is structured into several chapters, namely:

- Chapter 1: Introduction to Ipc Chemistry Review
- Chapter 2: Essential Elements of Ipc Chemistry Review
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- Chapter 5: Conclusion

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This book is crafted in an easy-to-understand language and is complemented by engaging illustrations. It is highly recommended for anyone seeking to gain a comprehensive understanding of Ipc Chemistry Review.

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**Table of Contents IPC Chemistry Review**

1. Understanding the eBook IPC Chemistry Review
  - The Rise of Digital Reading IPC Chemistry Review
  - Advantages of eBooks Over Traditional Books
2. Identifying IPC Chemistry Review
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an IPC Chemistry Review
  - User-Friendly Interface
4. Exploring eBook Recommendations from IPC Chemistry Review
  - Personalized Recommendations
  - IPC Chemistry Review User Reviews and Ratings
  - IPC Chemistry Review and Bestseller Lists
5. Accessing IPC Chemistry Review Free and Paid eBooks
  - IPC Chemistry Review Public Domain eBooks
  - IPC Chemistry Review eBook Subscription Services
  - IPC Chemistry Review Budget-Friendly Options
6. Navigating IPC Chemistry Review eBook Formats
  - ePub, PDF, MOBI, and More
  - IPC Chemistry Review Compatibility with Devices
  - IPC Chemistry Review Enhanced eBook Features
7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of IPC Chemistry Review
  - Highlighting and Note-Taking IPC Chemistry Review
  - Interactive Elements IPC Chemistry Review
8. Staying Engaged with IPC Chemistry Review

- Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers IPC Chemistry Review
9. Balancing eBooks and Physical Books IPC Chemistry Review
- Benefits of a Digital Library
  - Creating a Diverse Reading Collection IPC Chemistry Review
10. Overcoming Reading Challenges
- Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
11. Cultivating a Reading Routine IPC Chemistry Review
- Setting Reading Goals IPC Chemistry Review
  - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of IPC Chemistry Review
- Fact-Checking eBook Content of IPC Chemistry Review
  - Distinguishing Credible Sources
13. Promoting Lifelong Learning
- Utilizing eBooks for Skill Development
  - Exploring Educational eBooks
14. Embracing eBook Trends
- Integration of Multimedia Elements
  - Interactive and Gamified eBooks

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