

Isotopes Practice Set

- What does "mass" mean? \rightarrow mass is the amount of matter in an object. It is a scalar quantity and is measured in kilograms (kg).
- What does the number next to isotopes signify? \rightarrow The number next to an isotope symbol represents the mass number, which is the sum of protons and neutrons in the nucleus.
- How do you tell isotopes apart? \rightarrow Isotopes of an element have the same number of protons but different numbers of neutrons, resulting in different mass numbers.

For each of the following isotopes, write the number of protons, neutrons, and electrons.

	Chlorine-35	Chlorine-37
# of protons	17	17
# of neutrons	18	20
# of electrons	17	17

	Magnesium-24	Magnesium-26
# of protons	12	12
# of neutrons	12	14
# of electrons	12	12

	Carbon-12	Carbon-13
# of protons	6	6
# of neutrons	6	7
# of electrons	6	6

	Sulfur-32	Sulfur-34
# of protons	16	16
# of neutrons	16	18
# of electrons	16	16

Fill in the isotopes names including atomic numbers, including atomic numbers from the chart. Use your periodic table and the information provided.

	Isotope: $^{35}_{17}\text{Cl}$	Isotope: $^{37}_{17}\text{Cl}$
# of protons	17	17
# of neutrons	18	20
# of electrons	17	17

	Atom: $^{12}_{6}\text{C}$	Atom: $^{13}_{6}\text{C}$
# of protons	6	6
# of neutrons	6	7
# of electrons	6	6

	Carbon: $^{12}_{6}\text{C}$	Carbon: $^{13}_{6}\text{C}$
# of protons	6	6
# of neutrons	6	7
# of electrons	6	6

	Carbon: $^{12}_{6}\text{C}$	Carbon: $^{13}_{6}\text{C}$
# of protons	6	6
# of neutrons	6	7
# of electrons	6	6

Isotope Practice Set Answer Key

Wenbin Ji



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