

Name: KLDate: January 4, 2021

Student Exploration: Moles

Vocabulary: atomic mass, Avogadro constant, conversion factor, dimensional analysis, mole, molar mass, molecular mass, scientific notation, significant figures, unified atomic mass unit

Prior Knowledge Questions (Do these BEFORE using the Gizmo.)

- In the image to the right, note a dozen eggs, a dozen donuts and a dozen roses. How many of each item do you have? 12
- Would a dozen of each object have the same mass? No
- Suppose you have a dozen carbon atoms, a dozen gold atoms, and a dozen iron atoms. Even though you have the same number of each, would you expect them all to have the same mass? Explain.



No, I would not expect them all to have the same mass because each element has atoms that weigh differently. Their atoms differ in number of neutrons, protons and electrons, so the mass would be different for each dozen.

Gizmo Warm-up

When counting roses, eggs, or donuts, a dozen is a good unit to use. If you are counting atoms, however, a dozen is not much help. In the *Moles* Gizmo, you will learn about a unit used to count atoms.

On the **AVOGADRO CONSTANT** tab, place the copper (Cu) atom on the nano-balance on the left, which will show the average atomic mass of copper rather than the mass of a single copper atom.



- What is the average mass of a copper atom? 63.546 u

The unit "u" refers to **unified atomic mass units**. A single proton or neutron has a mass of approximately one atomic mass unit. (Officially, 1 u is one-twelfth the mass of a C-12 atom.)

- To gain an idea as to how many atoms are in a gram or so of copper, use the larger balance on the right. Press **Add atoms** to put a scoop of atoms in the weighing dish, and keep adding until the balance registers between 1 and 2 grams. If you don't seem to be making much progress, adjust the exponent using the slider, which will make the scoop size bigger.

How many atoms did you need to add? 1.00×10^{22} atoms

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Table of Contents Ionic Bonds Gizmo Key

1. Understanding the eBook Ionic Bonds Gizmo Key
 - The Rise of Digital Reading Ionic Bonds Gizmo Key
 - Advantages of eBooks Over Traditional Books
2. Identifying Ionic Bonds Gizmo Key
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Ionic Bonds Gizmo Key
 - User-Friendly Interface
4. Exploring eBook Recommendations from Ionic Bonds Gizmo Key
 - Personalized Recommendations
 - Ionic Bonds Gizmo Key User Reviews and Ratings
 - Ionic Bonds Gizmo Key and Bestseller Lists
5. Accessing Ionic Bonds Gizmo Key Free and Paid eBooks

- Ionic Bonds Gizmo Key Public Domain eBooks
 - Ionic Bonds Gizmo Key eBook Subscription Services
 - Ionic Bonds Gizmo Key Budget-Friendly Options
6. Navigating Ionic Bonds Gizmo Key eBook Formats
 - ePub, PDF, MOBI, and More
 - Ionic Bonds Gizmo Key Compatibility with Devices
 - Ionic Bonds Gizmo Key Enhanced eBook Features
 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Ionic Bonds Gizmo Key
 - Highlighting and Note-Taking Ionic Bonds Gizmo Key
 - Interactive Elements Ionic Bonds Gizmo Key
 8. Staying Engaged with Ionic Bonds Gizmo Key
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Ionic Bonds Gizmo Key
 9. Balancing eBooks and Physical Books Ionic Bonds Gizmo Key
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Ionic Bonds Gizmo Key
 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
 11. Cultivating a Reading Routine Ionic Bonds Gizmo Key
 - Setting Reading Goals Ionic Bonds Gizmo Key
 - Carving Out Dedicated Reading Time
 12. Sourcing Reliable Information of Ionic Bonds Gizmo Key
 - Fact-Checking eBook Content of Ionic Bonds Gizmo Key
 - Distinguishing Credible Sources
 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development

- Exploring Educational eBooks
14. Embracing eBook Trends
- Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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