

Making Copper—A Stoichiometry Experiment

ANALYSIS

1. How many grams of copper(II) chloride dihydrate ($\text{CuCl}_2 \cdot 2\text{H}_2\text{O}$) and aluminum were used in the experiment? Show calculations where relevant.

$$\begin{array}{l}
 \text{3 moles} \\
 3 (170.5 \text{ g}) \\
 = 511.5 \text{ g}
 \end{array}
 \qquad
 \begin{array}{l}
 \text{2 moles} \\
 2 (27 \text{ g}) \\
 = 54 \text{ g}
 \end{array}
 \qquad
 \begin{array}{l}
 \text{3 moles} \\
 3 (63.5 \text{ g}) \\
 = 190.5 \text{ g}
 \end{array}$$

Answer: 511.5 g

Answer: 54 g

2. Balance the equation for the reaction between $\text{CuCl}_2 \cdot 2\text{H}_2\text{O}$ and Al.



3. Use the masses of $\text{CuCl}_2 \cdot 2\text{H}_2\text{O}$ and Al from question #1 to calculate the mass of Cu (theoretical yield) produced. (Note: When finding the molar mass of $\text{CuCl}_2 \cdot 2\text{H}_2\text{O}$ be sure to add the mass of 2 moles of water to the molar mass of CuCl_2).

$$\frac{190.5}{511.5} \times 100 = 37.24\%$$

Answer: 37.24%

Fe And Copper Chloride Dihydrate Stoichiometry

J Dewey



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