

STATES OF MATTER QUESTIONS TEST PAPER- I

1. Define Boyle's law and write its mathematical expression.
Express it graphically.
2. State Charles Law. Define absolute zero temperature. Write the characteristics and express it graphically.
3. Define an ideal gas.
4. What is aqueous tension?
5. What is the value of R at STP?
6. Molecule A is twice as heavy as the molecule B. which of these has higher kinetic energy at any temperature?
7. Write Van der waal's equation for n moles of a gas.
8. Under what conditions do real gases tend to show ideal gas behaviour?
9. Deduce the relation $pV = nRT$ where R is a constant called universal
10. Mention the two assumptions of kinetic theory of gases that do not hold good.
11. What is the relationship between the density and molar mass of a gaseous substance? Derive it.
12. State Dalton's law of partial pressure. How will you calculate the partial pressure of gas using Dalton's Law of partial pressures?
13. What do you understand by the terms, Isotherms, Isochores and Isobars?
14. State the fundamental assumptions of the kinetic theory of gases
15. What property of molecules of real gases is indicated by van der waal's constant 'a'?
Ans: **intermolecular force**
16. Match the following gas laws with the equation representing them.

Matter And Material Question Paper Grade 1mamo

Julia Schneider



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