

Stoichiometry Practice 1.1 - Answer Key

Objectives: - Use stoichiometry to calculate the mass of any substance involved in a chemical reaction.

Directions: Answer the following stoichiometry problems and show your work.

Mass A → Mole A → Mole B:

1. When calcium hydroxide and phosphoric acid (H_3PO_4) are combined, calcium phosphate and water are produced. If the reaction starts with 10.0 grams of phosphoric acid, how many moles of water are produced?



$$10.0 \text{ g} = \left(\frac{2 \text{ mol H}_3\text{PO}_4}{326.06 \text{ g H}_3\text{PO}_4} \right) \left(\frac{6 \text{ mol H}_2\text{O}}{2 \text{ mol H}_3\text{PO}_4} \right) = \boxed{0.368 \text{ mol H}_2\text{O}}$$

Mole A → Mole B → Mass B:

2. How many grams of oxygen are needed in the combustion 1.07 mol of methyl alcohol (CH_3OH)?



$$1.07 \text{ mol} = \left(\frac{3 \text{ mol O}_2}{2 \text{ mol CH}_3\text{OH}} \right) \left(\frac{32.00 \text{ g O}_2}{1 \text{ mol O}_2} \right) = \boxed{51.4 \text{ g}}$$

Mass A → Mole A → Mole B → Mass B:

3. How many grams of copper are produced when 3.46 grams of aluminum react with excess copper (II) acetate solution? $\boxed{13.2 \text{ g}}$



$$3.46 \text{ g} = \left(\frac{2 \text{ mol Al}}{26.98 \text{ g Al}} \right) \left(\frac{3 \text{ mol Cu}}{2 \text{ mol Al}} \right) \left(\frac{63.546 \text{ g Cu}}{1 \text{ mol Cu}} \right) = \boxed{13.2 \text{ g Cu}}$$

4. Lithium sulfate and barium chloride react to produce barium sulfate and lithium chloride. How many grams of barium chloride are needed to react with 17.5 g of lithium sulfate?



$$17.5 \text{ g} = \left(\frac{1 \text{ mol Li}_2\text{SO}_4}{148.97 \text{ g Li}_2\text{SO}_4} \right) \left(\frac{1 \text{ mol BaCl}_2}{1 \text{ mol Li}_2\text{SO}_4} \right) \left(\frac{208.23 \text{ g BaCl}_2}{1 \text{ mol BaCl}_2} \right) = \boxed{23.8 \text{ g}}$$

Mass Mass Stoichiometry Practice Key

**Donald J. Wink, Sharon Fetzer-
Gislason, Sheila McNicholas**

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Table of Contents Mass Mass Stoichiometry Practice Key

1. Understanding the eBook Mass Mass Stoichiometry Practice Key
 - The Rise of Digital Reading Mass Mass Stoichiometry Practice Key
 - Advantages of eBooks Over Traditional Books
2. Identifying Mass Mass Stoichiometry Practice Key
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Mass Mass Stoichiometry Practice Key
 - User-Friendly Interface
4. Exploring eBook Recommendations from Mass Mass Stoichiometry Practice Key
 - Personalized Recommendations
 - Mass Mass Stoichiometry Practice Key User Reviews and Ratings
 - Mass Mass Stoichiometry Practice Key and Bestseller Lists

5. Accessing Mass Mass Stoichiometry Practice Key Free and Paid eBooks
 - Mass Mass Stoichiometry Practice Key Public Domain eBooks
 - Mass Mass Stoichiometry Practice Key eBook Subscription Services
 - Mass Mass Stoichiometry Practice Key Budget-Friendly Options
6. Navigating Mass Mass Stoichiometry Practice Key eBook Formats
 - ePub, PDF, MOBI, and More
 - Mass Mass Stoichiometry Practice Key Compatibility with Devices
 - Mass Mass Stoichiometry Practice Key Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Mass Mass Stoichiometry Practice Key
 - Highlighting and Note-Taking Mass Mass Stoichiometry Practice Key
 - Interactive Elements Mass Mass Stoichiometry Practice Key
8. Staying Engaged with Mass Mass Stoichiometry Practice Key
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Mass Mass Stoichiometry Practice Key
9. Balancing eBooks and Physical Books Mass Mass Stoichiometry Practice Key
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Mass Mass Stoichiometry Practice Key
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Mass Mass Stoichiometry Practice Key
 - Setting Reading Goals Mass Mass Stoichiometry Practice Key
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Mass Mass Stoichiometry Practice Key
 - Fact-Checking eBook Content of Mass Mass Stoichiometry Practice Key
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
- Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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