

## IB Chemistry – HL

### Atomic structure

1. What is the electron configuration for an atom with  $Z = 22$ ?

- A.  $1s^2 2s^2 2p^6 3s^2 3p^6 3d^4$
- B.  $1s^2 2s^2 2p^6 3s^2 3p^6 4s^2 4p^2$
- C.  $1s^2 2s^2 2p^6 3s^2 3p^6 3d^2 4p^2$
- D.  $1s^2 2s^2 2p^6 3s^2 3p^6 4s^2 3d^2$

(Total 1 mark)

2. What is the total number of p orbitals containing one or more electrons in germanium (atomic number 32)?

- A. 2
- B. 3
- C. 5
- D. 8

(Total 1 mark)

3. How many electrons are there in all the d orbitals in an atom of xenon?

- A. 10
- B. 18
- C. 20
- D. 36

(Total 1 mark)

4. Which is correct about the element tin (Sn) ( $Z = 50$ )?

Number of electrons in highest main energy level	Number of main energy levels containing electrons	
4	4	A.
14	4	B.
4	5	C.
14	5	D.

(Total 1 mark)

# May 2014 Paper 1b Chem HL

**J Spring**



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