

## Student Exploration: Covalent Bonds

<p><b>Activity B:</b> <b>Building molecules</b></p>	<p><u>Get the Gizmo ready:</u></p> <ul style="list-style-type: none"> <li>• Click <b>Reset</b>.</li> <li>• Turn off <b>Show Lewis diagram</b>.</li> <li>• Select <b>Oxygen</b>.</li> </ul>	
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**Question:** How do atoms share more than one pair of electrons?

1. **Observe:** Like fluorine and most other elements, oxygen atoms are most stable with a full complement of eight valence electrons.

A. How many valence electrons does each oxygen atom have now? 6

B. How many more electrons does each oxygen atom need to be stable? 2

2. **Form a bond:** Drag electrons back and forth until the molecule of oxygen ( $O_2$ ) is stable. Click **Check** to confirm your molecule is stable. Take a snapshot and paste the image into your document (don't forget to label it " $O_2$ ").

How many *pairs* of shared electrons are there in a stable molecule of oxygen? 1

3. **Draw a diagram:** Draw a Lewis diagram of the oxygen molecule in the space below at left. To check your work, turn on **Show Lewis diagram**. Draw the correct diagram on the right.

Practice diagram:  $\cdot\ddot{O} = \ddot{O}\cdot$

Actual:  $\cdot\ddot{O} = \ddot{O}\cdot$

4. **Practice:** Create covalent bonds and stable molecules for the remaining substances. Take a snapshot of each completed molecule and add it to your document. Draw Lewis diagrams for each one. (As above, draw the diagram on your own before checking your work.)

Nitrogen



Ammonia



Chlorine



Methane



Water



Formaldehyde



(Activity B continued on next page)

# Explore Learning Covalent Bonds Answer Key

**EW Minium**

A decorative graphic element consisting of a light blue horizontal bar with a rounded right end, and a red-to-white gradient semi-circle overlapping the right side of the bar.

## **Explore Learning Covalent Bonds Answer Key:**

## Adopting the Song of Expression: An Emotional Symphony within **Explore Learning Covalent Bonds Answer Key**

In a global eaten by monitors and the ceaseless chatter of instantaneous communication, the melodic elegance and psychological symphony created by the published term usually disappear in to the back ground, eclipsed by the constant sound and disruptions that permeate our lives. However, situated within the pages of **Explore Learning Covalent Bonds Answer Key** a stunning literary value brimming with organic thoughts, lies an immersive symphony waiting to be embraced. Constructed by a masterful musician of language, this fascinating masterpiece conducts visitors on a psychological trip, skillfully unraveling the concealed songs and profound affect resonating within each carefully constructed phrase. Within the depths of the moving analysis, we will explore the book is central harmonies, analyze its enthralling writing style, and surrender ourselves to the profound resonance that echoes in the depths of readers souls.

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