

Intermolecular Forces

Intermolecular Forces:

- Forces between molecules
- Responsible for the state of matter: solid, liquid, or gas
- affect the melting and boiling points of compounds as well as the solubility of one substance in another
- weaker than covalent bonds within molecules, since molecular compounds melt easily (melting doesn't break the bond between atoms, but between the molecules themselves!)
 - o example: melting ice to water to steam will break apart the intermolecular forces and change state, but not decompose into H_2 and O_2

Intramolecular Force (Force-Force attraction):

- Forces within molecules
- the attractive force between the valence electrons of one atom and the nucleus of another atom would be considered intermolecular forces
- Covalent bonds are examples of intramolecular forces

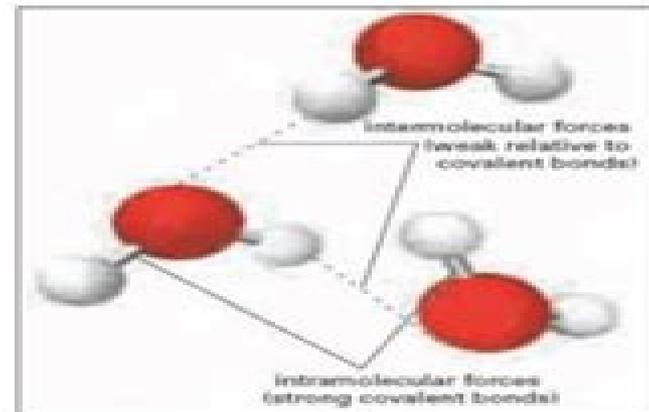
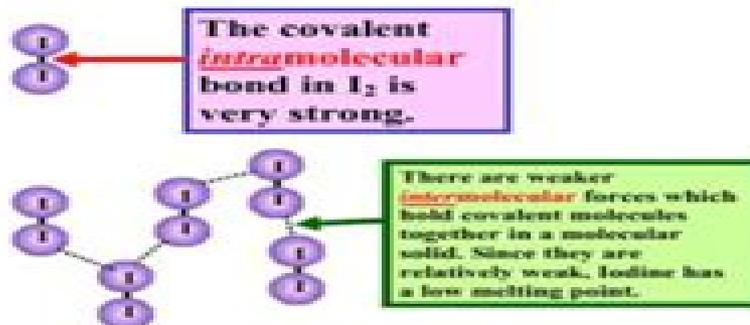


Figure 3.20 Strong intramolecular forces (covalent bonds) hold the atoms in molecules together. Relatively weak intermolecular forces act between molecules.

Exploring Intermolecular Forces Lab Answers

RJ Shavelson



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