

Name: _____

Unit 6: Exponents & Exponential Functions



Date: _____

Bell: _____

Homework 4: Negative Exponents

Directions: Simplify the following monomials. Your answer should contain positive exponents only!

<p>1. x^{-7}</p> $\frac{1}{x^7}$	<p>2. $5x^2y^{-3}$</p> $\frac{5x^2}{y^3}$	<p>3. $-4a^{-2}b^{-2}$</p> $\frac{-4}{a^2b^2}$
<p>4. $(ab^2)^{-4}$</p> $a^{-4}b^{-8}$ $= \frac{1}{a^4b^8}$	<p>5. $-8(x^3y^4)^{-5}$</p> $-8x^{15}y^{-20}$ $= \frac{-8x^{15}}{y^{20}}$	<p>6. $(3x)^{-3}$</p> $\frac{1}{27}x^{-3}$ $= \frac{1}{27x^3}$
<p>7. $(a^2)(a^{-3})$</p> a^{-5} $= \frac{1}{a^5}$	<p>8. $(x^{-2})^{-3}(-2y^5)^4$</p> $x^{-9} \cdot 16y^{20}$ $= \frac{16y^{20}}{x^9}$	<p>9. $(a^2b^3)^{-2}(a^5b^4)^{-3}$</p> $a^{-4}b^{-6} \cdot a^{-15}b^{-12}$ $= \frac{1}{a^{19}b^{18}}$
<p>10. $(2x^4)^{-5}$</p> $\frac{1}{32x^{20}}$	<p>11. $(-11x^2)^{-2}$</p> $\frac{1}{121x^4}$	<p>12. $(x^3y^6)^2 + (x^2y^4)^{-2}$</p> $x^6y^{12} + x^{-4}y^{-8}$ $\frac{x^6y^{12}}{x^4y^8} + \frac{1}{x^4y^8}$
<p>$= \frac{x^2y^4}{1}$</p> $\frac{x^2y^4}{1}$	<p>$= \frac{1}{121x^4}$</p> $\frac{1}{121x^4}$	<p>$= \frac{x^2y^4}{1} + \frac{1}{x^4y^8}$</p> $\frac{x^2y^4}{1} + \frac{1}{x^4y^8}$
<p>$= \frac{x^2y^4}{1}$</p> $\frac{x^2y^4}{1}$	<p>$= \frac{1}{121x^4}$</p> $\frac{1}{121x^4}$	<p>$= \frac{x^2y^4}{1} + \frac{1}{x^4y^8}$</p> $\frac{x^2y^4}{1} + \frac{1}{x^4y^8}$
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web the pcl₃ lewis structure has the typical case of phosphorus p in the center with 3 bonds to 3 other atoms phosphorus is from the same column as nitrogen in the periodic table meaning that p and n generally have the same bonding structure note the lone pair dots without bonds on top of p just like for n in the previous example for nh₃ chlorine cl

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web step 1 figure out how many electrons the molecule must have based on the number of valence electrons in each atom when drawing the structure of an ion be sure to add subtract electrons to account for the charge step 2 connect the atoms to each other with single bonds to form a skeleton structure

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web lewis structures a lewis structure is a representation of covalent molecules or polyatomic ions where all the valence

electrons are shown distributed about the bonded atoms as either shared electron pairs bond pairs or unshared electron pairs lone pairs a shared pair of electrons is represented as a short line a single bond

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2 total number of valence electrons 24 cad engineered 3d sketch model show dipole arrows lewis structure is there a polar bond in this molecule yes or no vsepr shape name trigonal planar bond angles 120

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web jun 8 2022 lewis structure in chemistry lewis structures also called electron dot structures or electron dot diagrams are diagrams that show the bonding between atoms of a molecule and the lone pairs of

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web this exercise required us to take the formulas in the data table and calculate the number of valence electrons using the periodic table draw its lewis structure on paper build the molecular model and record a photo list the number of bonds and lone pairs surrounding the central atom and lastly identify the structure geometry