

## **GRAVIMETRIC ANALYSIS OF A CHLORIDE SALT**

*Date submitted: 14/03/2025*

*By: ()*

*Lab section:*

*TA*

**Purpose:** The purpose of this experience is to determine the quantity of the chloride content in an unknown soluble salt.

### **Theory:**

The assignment in this lab was to find the concentration of chloride ion in the given unknown salt sample. Gravimetric analysis is a quantitative analytical technique where the amount of an analyte is determined by measuring the mass of a solid precipitate.

**Precipitation Reaction of Chloride Ions with Silver Ions:** The reaction between chloride ions and silver ions forms a solid precipitate of silver chloride:



So,  $n(\text{AgCl})$  is equal to  $n(\text{Cl})$

### **Solubility of AgCl (s) in water:**

According to the lab manual section "Theory"  $\text{AgCl}(\text{s})$  is very insoluble, which lead to the completion of the reaction.

$$K_{sp} = [\text{Ag}^+][\text{Cl}^-] = 1.6 \times 10^{-10}$$

### **Calculation for the amount of precipitate lost by washing:**

Even though,  $\text{AgCl}$  is insoluble during the washing process some of the precipitate was lost.

Therefore:  $K_{sp} = [\text{Ag}(\text{aq})^+] \cdot [\text{Cl}(\text{aq})^-] = 1.6 \times 10^{-10}$

$$[x] \cdot [x] = 1.6 \times 10^{-10}$$

$$x = 1.3 \times 10^{-5}$$

-  $x$  is equal to  $1.3 \times 10^{-5}$  moles per litre

### **Precipitation is performed in an acidic environment:**

# Gravimetric Analysis Of Chloride Salt Lab Report

**Y Pai**



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