
Bond Energy Study Guide (Chemistry 2.2)

1. Why atoms form bonds

Atoms form bonds to reach a lower, more stable energy state.

- Each atom has **protons (+)** in its nucleus and **electrons (-)** around it.
- **Coulombic attraction** is the force between opposite charges. When atoms get close, their nuclei attract the other atom's electrons.
- This attraction lowers potential energy — that's why bonds form.

When the bond forms:

- **Energy is released** → the system becomes more stable.
- **Bond making = exothermic process.**

When a bond breaks:

- **Energy is absorbed** → you must overcome that attraction.
- **Bond breaking = endothermic process.**

2. Endothermic vs. Exothermic

| Type of Process | Energy Flow | What Happens | ΔH (Enthalpy Change) Sign |
|-----------------|-----------------|------------------|-----------------------------------|
| Endothermic | Energy absorbed | Bonds are broken | Positive (+) |
| Exothermic | Energy released | Bonds are formed | Negative (-) |

Example:

- When hydrogen and chlorine form HCl, the reaction releases energy — **exothermic**.
- When you separate HCl into H and Cl atoms, you need energy — **endothermic**.

3. Net Enthalpy Change (ΔH)

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