

[4236] - 301
M.Sc. - II
ELECTRONIC SCIENCE
EL3 UT05 : Embedded Systems
(2008 Pattern) (Sem. - III)

Time : 3 Hours]

[Max. Marks : 80

Instructions to the candidates:

- 1) All questions are compulsory.*
- 2) Figures to the right indicate full marks.*
- 3) Draw neat diagrams wherever necessary.*

Q1) Attempt any two of the following :

[2 × 8 = 16]

- a) Draw programming model of 8051 showing different registers without their addresses. Explain TCON and PSW registers of 8051.
- b) State the features of I²C bus and explain with waveforms following conditions.
 - i) Start
 - ii) Stop
 - iii) ACKExplain data format used in I²C bus standard.
- c) Write a program in C for single digit decimal calculator for addition and subtraction. Read input from keyboard and display result on LCD. (8051 microcontroller)

Q2) Attempt any two of the following :

[2 × 8 = 16]

- a) Write function of EA and ALE pins of 8051. Draw and explain the circuit for interfacing of external 8k × 8bit EPROM and 8k × 8bit RAM to 8051 using suitable 3:8 decoder. Show all control lines and memory map.
- b) Draw schematic diagram of 8051 based target board with serial port interface to computer. List the components required to design target board. Explain how you will test crystal and ports. Explain reset circuit used in 8051 based system.
- c) Draw circuit diagram to interface two stepper motors to port 1 of 8051. Write C program to rotate stepper motors one in clockwise and other in anticlockwise direction continuously.

P.T.O.

Embedded Systems Question Paper

Lingjun Ying



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web solution 1 the two half reactions balanced as if in acidic solution $2\text{CN}^- + \text{Au} + \text{CN}^- + 2\text{e}^- + 2\text{H}^+ + \text{O}_2 \rightarrow \text{H}_2\text{O}_2 + 2\text{e}^-$ make electrons equal convert to basic solution $4\text{CN}^- + 2\text{Au} + 2\text{OH}^- + 2\text{e}^-$ multiplied by a factor of 2 $2\text{e}^- + 2\text{H}_2\text{O} + \text{O}_2 \rightarrow \text{H}_2\text{O}_2 + 2\text{OH}^- + 3\text{e}^-$ the final answer

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web the half equation method separates the oxidation and reduction of a redox reaction in half reactions overall scheme for the half reaction method step 1 split reaction into half reactions reduction and oxidation step 2 balance the charge or oxidation number with electrons step 3 balance ce o by adding ce h₂o

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web we can balance it for charge by adding two electrons to the right side of the equation so that the net charge on each side is 0 oxidation $\text{Ni} \rightarrow \text{Ni}^{2+} + 2\text{e}^-$ now that the oxidation half reaction is balanced it tells us that two electrons are produced for every atom of nickel oxidized but where do those electrons go

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web balancing redox reactions in acidic solution problems 1 10 problem 1 $\text{Cr}_2\text{O}_7^{2-} + \text{Fe}^{2+} + \text{Cr}^{3+} + \text{Fe}^{3+}$ solution 1 balanced half reactions $6\text{e}^- + 14\text{H}^+ + \text{Cr}_2\text{O}_7^{2-} \rightarrow 2\text{Cr}^{3+} + 7\text{H}_2\text{O}$ $\text{Fe}^{2+} \rightarrow \text{Fe}^{3+} + \text{e}^-$ 2 equalize the electrons $6\text{e}^- + 14\text{H}^+ + \text{Cr}_2\text{O}_7^{2-} \rightarrow 2\text{Cr}^{3+} + 7\text{H}_2\text{O}$ $6\text{Fe}^{2+} \rightarrow 6\text{Fe}^{3+} + 6\text{e}^-$ multiplied by a factor of 6

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reaction in acid solution using the half reaction method $8 \text{H}_2\text{O}_2 + \text{Cr}_2\text{O}_7^{2-} + 2 \text{H}^+ + 10 \text{e}^- \rightarrow 8 \text{H}_2\text{O} + \text{Cr}^{3+} + 11 \text{H}^+$
as $\text{NO}_3^- + 3 \text{e}^- + 4 \text{H}^+ \rightarrow \text{H}_2\text{SO}_4$

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web in the redox equation below that occurs in our stomach which of the following pairs identifies the reducing and oxidizing agents respectively $\text{C}_6\text{H}_8\text{O}_6 + 2 \text{HNO}_2 + 2 \text{C}_6\text{H}_6\text{O}_6 + 2 \text{H}_2\text{O} + 2 \text{NO}$ choose 1 answer $\text{C}_6\text{H}_8\text{O}_6$ and NO_2 a $\text{C}_6\text{H}_8\text{O}_6$ and NO_2 b NO_2 and $\text{C}_6\text{H}_8\text{O}_6$ c H_2O and $\text{C}_6\text{H}_8\text{O}_6$ d $\text{C}_6\text{H}_8\text{O}_6$ and H_2O

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