

Chemistry Review Packet

Name: _____

Date: _____

1. Consider the following reaction:



When 2.0 moles of hydrogen are reacted with 48.0 grams of oxygen, how many grams of water are formed?

- A. 18.0 grams B. 36.0 grams C. 54.0 grams D. 11.0 grams E. 90.0 grams

2. Which of the following 1.0 M aqueous solutions has the lowest pH?

- A. HNO_3 B. NaCl C. HNO_2 D. NaOH E. HCO_2H

3. Which of the following substances cannot be broken down into simpler substances by physical or chemical means?

- A. Wood B. Baking soda C. Natural gas D. Neon gas E. Water

4. What is the approximate percent composition by mass of hydrogen in the compound ethanol, $\text{C}_2\text{H}_6\text{O}$ (molecular weight, in grams: 46.10 or 46.08 or 46.0)?

- A. 4.3% B. 12.0% C. 29.0% D. 34.0% E. 20.0%

5. The empirical formula of a certain hydrocarbon is C_2H_4 , and it has a molecular weight equal to 80 g/mol. Based on this information, the molecular formula of the hydrocarbon is

- A. C_2H_4 B. C_4H_8 C. C_6H_{12} D. C_8H_{16} E. $\text{C}_{10}\text{H}_{20}$

6. Which of the following aqueous solutions will have the highest normal boiling point?

- A. 0.20 M CaCl_2 B. 0.20 M NaCl C. 0.5 M KNO_3 D. 1.0 M $\text{C}_2\text{H}_6\text{O}_2$ E. 0.20 M MgCl_2

7. Consider the reaction



How is this reaction best classified?

- A. Precipitation B. Redox reaction C. Single replacement D. Combustion E. Synthesis

8. The following reaction takes place under constant temperature and pressure:



When 100 L of hydrogen gas reacts completely with excess oxygen, how many liters of gas remain? (Note: All gases are at the same temperature and pressure.)

- A. 100 L B. 150 L C. 400 L D. 100 L E. 200 L

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Robert R. Harr



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