

### Practice problems for Gravimetric analysis and gravimetric factor ratio Answers

1. Determine the gravimetric factor for the following sought analytes and weighed final products:

(a) Weighed final product:  $\text{CaC}_2\text{O}_4$

Sought analyte:  $\text{CaCO}_3$

$$\text{MW}(\text{CaC}_2\text{O}_4) = 40.08 \text{ g/mol} + 2(12.01 \text{ g/mol}) + 4(16.00 \text{ g/mol}) = 128.10 \text{ g/mol}$$

$$\text{MW}(\text{CaCO}_3) = 40.08 \text{ g/mol} + 12.01 \text{ g/mol} + 3(16.00 \text{ g/mol}) = 100.09 \text{ g/mol}$$

$$GF = \frac{1 \times \text{MW of CaCO}_3}{1 \times \text{MW of CaC}_2\text{O}_4} = \frac{100.09}{128.10} = 0.78134$$

(b) Weighed final product:  $\text{PbI}_2$

Sought analyte:  $\text{NaIO}_3$

$$\text{MW}(\text{PbI}_2) = 207.2 \text{ g/mol} + 2(126.90 \text{ g/mol}) = 461.0 \text{ g/mol}$$

$$\text{MW}(\text{NaIO}_3) = 22.99 \text{ g/mol} + 126.90 \text{ g/mol} + 3(16.00 \text{ g/mol}) = 197.89 \text{ g/mol}$$

$$GF = \frac{2 \times \text{MW of NaIO}_3}{1 \times \text{MW of PbI}_2} = \frac{2(197.89)}{461.0} = 0.8585$$

(c) Weighed final product:  $\text{B}_2\text{O}_3$

Sought analyte:  $\text{Na}_2\text{B}_4\text{O}_7$

$$\text{MW}(\text{B}_2\text{O}_3) = 2(10.81 \text{ g/mol}) + 3(16.00 \text{ g/mol}) = 69.62 \text{ g/mol}$$

$$\text{MW}(\text{Na}_2\text{B}_4\text{O}_7) = 2(22.99 \text{ g/mol}) + 4(10.81) \text{ g/mol} + 7(16.00 \text{ g/mol}) = 201.22 \text{ g/mol}$$

$$GF = \frac{1 \times \text{MW of Na}_2\text{B}_4\text{O}_7}{2 \times \text{MW of B}_2\text{O}_3} = \frac{201.22}{2(69.62)} = 1.445$$

(d) Weighed final product:  $\text{Hg}_2\text{I}_2$

Sought analyte:  $\text{HgO}$

$$\text{MW}(\text{Hg}_2\text{I}_2) = 2(200.59 \text{ g/mol}) + 2(126.90 \text{ g/mol}) = 654.98 \text{ g/mol}$$

$$\text{MW}(\text{HgO}) = 200.59 \text{ g/mol} + 16.00 \text{ g/mol} = 216.59 \text{ g/mol}$$

$$GF = \frac{2 \times \text{MW of HgO}}{1 \times \text{MW of Hg}_2\text{I}_2} = \frac{2(216.59)}{654.98} = 0.66136$$

2. The hydrogen sulphide ( $\text{H}_2\text{S}$ ) in a 50.0 g sample of crude petroleum was removed by distillation and collected in a solution of  $\text{CdCl}_2$ . The precipitated  $\text{CdS}$  was then filtered, washed, and ignited to  $\text{CdSO}_4$ . Calculate the percentage of  $\text{H}_2\text{S}$  in the sample if 0.108 g of  $\text{CdSO}_4$  was recovered.

$$\text{MW}(\text{H}_2\text{S}) = 2(1.01) \text{ g/mol} + 32.07 \text{ g/mol} = 34.09 \text{ g/mol}$$

$$\text{MW}(\text{CdSO}_4) = 112.41 \text{ g/mol} + 32.07 \text{ g/mol} + 4(16.00 \text{ g/mol}) = 208.48 \text{ g/mol}$$

$$GF = \frac{1 \times \text{MW of H}_2\text{S}}{1 \times \text{MW of CdSO}_4} = \frac{34.09}{208.48} = 0.1635$$

$$\% \text{H}_2\text{S} = [(0.108 \text{ g}) \times (0.1635)/50.0 \text{ g} \times 100] = 0.0353\%$$

# Gravimetric Analysis Problems With Answers

**Wenbin Ji**



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