

Equilibrium and Concentration

Answer Key

Vocabulary: chemical equilibrium, concentration, equilibrium, equilibrium constant, reaction quotient, reversible reaction

Prior Knowledge Questions (Do these BEFORE using the Gizmo.)

[Note: The purpose of these questions is to activate prior knowledge and get students thinking. Students are not expected to know the answers to the Prior Knowledge Questions.]

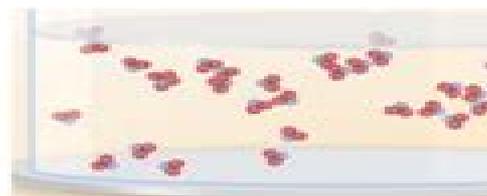
Gary has \$5,000 in his bank account and earns a modest salary. Every month he pays for rent, food, utilities, and entertainment.

- How will Gary's account change if he saves more than he spends? *It will increase.*
- How will Gary's account change if he spends more than he saves? *It will decrease.*
- What happens if Gary spends exactly as much as he saves? *It will stay the same.*

Gizmo Warm-up

If Gary spends exactly as much as he earns, his savings will be in **equilibrium**. Equilibrium occurs when two opposing processes occur at the same rate, leading to no net change. In the *Equilibrium and Concentration Gizmo™*, you will investigate how equilibrium can occur in chemical reactions.

To begin, check that **Reaction 1** is selected. Set **Moles NO₂** to 8 and **Moles N₂O₄** to 0.



- Click **Play** (▶) and observe the colliding molecules. What do you notice?

Sample answer: The molecules move and collide. Occasionally there is a blue or red flash.

In the Gizmo, a blue flash appears every time two reactants combine to form a product. A red flash appears every time a product dissociates into reactants.

- Click **Reset** (↺), and set **Moles NO₂** to 0 and **Moles N₂O₄** to 8. Click **Play**.

What do you notice now? *Sample answer: At first, there are more red flashes than blue flashes.*

- When a reaction can proceed in either direction, it is a **reversible reaction**. Based on what you have observed, is the synthesis of NO₂ into N₂O₄ a reversible reaction? Explain.

I know the synthesis of NO₂ into N₂O₄ is a reversible reaction because both the forward reaction (blue flashes) and the reverse reaction (red flashes) are observed.

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RS Peters



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