

## EXTRA PRACTICE PROBLEMS FOR MODULE #6

1. Nitric acid is manufactured with the following reaction:



How many grams of nitrogen dioxide are required to produce  $5.89 \times 10^3$  kg  $\text{HNO}_3$  in excess water?

2. Tungsten metal, W, is produced by the following reaction:



How many grams of tungsten can be made from  $1.0 \times 10^6$  g of hydrogen in excess  $\text{WO}_3$ ?

3. How many grams of  $\text{WO}_3$  would be used in the above reaction under the conditions specified in #2?  
4. Carbon tetrachloride is made according to the following reaction:



How many grams of  $\text{CS}_2$  are needed to make 16 kg of  $\text{CCl}_4$ ?

5. To make a phosphorus fertilizer ( $\text{CaH}_4\text{P}_2\text{O}_8$ ), agricultural companies use the following reaction:



If  $2.50 \times 10^3$  grams of  $\text{Ca}_3\text{P}_2\text{O}_8$  are reacted with excess  $\text{H}_2\text{SO}_4$  and  $\text{H}_2\text{O}$ , how many grams of fertilizer can be made?

6. In the following reaction:



What mass of Mg must you use to make  $5.6 \times 10^4$  g of  $\text{Mg(OH)}_2$ ?

7. Which of the following are empirical formulas?

a.  $\text{SO}_3$

b.  $\text{K}_2\text{S}_2\text{O}_3$

c.  $\text{C}_6\text{H}_{12}\text{O}_6$

d.  $\text{P}_2\text{H}_4$

8. Decomposition of 100.0 g of a dark green powder yields 39.7 g K, 27.9 g Mn, and an unknown amount of oxygen. What is the empirical formula?

9. The mass of the molecule above is 197.1 amu. What is the molecular formula?

10. A compound has an empirical formula of  $\text{CH}_2\text{O}$ . If its molecular mass is 60.0 amu, what is its molecular formula?

# Exploring Creation With Chemistry Module Answers

**Camilla Rothe**



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