

Lesson 3: Limiting reactants

- Reactions stop when one of the reactants is used up.

Limiting reactant - is used completely in the reaction; determines the amount of product formed

Excess reactant - leftover once the reaction stops.

- To calculate which reactant is limiting, you will compare the amount of each reactant to a single product (the true amount of product formed came from the limiting)



If 200.0g of sulfur reacts with 100.0g of chlorine, what mass of disulfur dichloride is produced? which is the limiting reactant?

given = 200.0g S_8 ; 100.0g Cl_2

unknown = ? g S_2Cl_2 ; ? limiting reactant

$$200.0g S_8 \times \frac{1 \text{ mol } S_8}{256.528g S_8} \times \frac{4 \text{ mol } S_2Cl_2}{1 \text{ mol } S_8} \times \frac{135.038g S_2Cl_2}{1 \text{ mol } S_2Cl_2} = 421g S_2Cl_2$$

$$100.0g Cl_2 \times \frac{1 \text{ mol } Cl_2}{70.906g Cl_2} \times \frac{4 \text{ mol } S_2Cl_2}{4 \text{ mol } Cl_2} \times \frac{135.038g S_2Cl_2}{1 \text{ mol } S_2Cl_2} = \boxed{190g S_2Cl_2}$$

Chlorine is the limiting reactant



If 100.0g Na and 100.0g Fe_2O_3 react, what mass of Iron is produced? which is the limiting reactant?

given = 100.0g Na; 100.0g Fe_2O_3

unknown = ? g Fe; ? limiting reactant

$$100.0g Na \times \frac{1 \text{ mol } Na}{22.990g Na} \times \frac{2 \text{ mol } Fe}{6 \text{ mol } Na} \times \frac{55.845g Fe}{1 \text{ mol } Fe} = 81.0g Fe$$

$$100.0g Fe_2O_3 \times \frac{1 \text{ mol } Fe_2O_3}{159.687g Fe_2O_3} \times \frac{2 \text{ mol } Fe}{1 \text{ mol } Fe_2O_3} \times \frac{55.845g Fe}{1 \text{ mol } Fe} = \boxed{70.0g Fe}$$

L.R. = Fe_2O_3

Limiting Reactants Study Guide

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