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# LIMITING REACTANT LAB

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**MATERIALS:** 250 mL beaker, 100 mL graduated cylinder, spoon or stirring rod, and a watch glass.

**SAFETY PRECAUTIONS:** Wear goggles during this experiment. Avoid skin contact with copper (II) chloride dihydrate. Wash your hands before leaving the laboratory.

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## PRELAB QUESTION:

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Aluminum metal reacts with aqueous copper (II) chloride dihydrate ( $\text{CuCl}_2 \cdot 2 \text{H}_2\text{O}$ ) to form copper metal, aqueous aluminum chloride, and water. Write the balanced equation for this reaction, including state symbols.

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## PROCEDURE:

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### DAY 1

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1. Weigh out the copper (II) chloride dihydrate. Record the true mass of the hydrated compound used.
2. Using the chart below, select your group number and the approximate masses of aluminum metal and copper (II) chloride dihydrate you'll be using in this activity:

Groups	Mass of Al (g)	Mass of $\text{CuCl}_2 \cdot 2 \text{H}_2\text{O}$ (g)
1 & 7 & 13	1.00	2.00
2 & 8 & 14	0.50	2.00
3 & 9	0.50	4.00
4 & 10	0.50	6.00
5 & 11	0.25	5.00
6 & 12	0.25	4.00

3. Using a graduated cylinder, measure water (tap water is fine) and add it to your LABELED beaker containing the copper (II) chloride dihydrate. Stir until the salt dissolves completely. Keep adding water (maximum of 100 mL) until all salt is dissolved – record the amount of water used. Note the color of the solution.
4. Weigh out the appropriate mass of aluminum foil, record that mass, then crumple it up into a few small pieces (NOT TIGHTLY CRUMPLED) and add it to the copper (II) chloride dihydrate solution. Stir the solution gently and then allow it to sit undisturbed for 5-10 minutes. Write observations.

# Limiting Reagent Lab Answer Key

**Mark Steyvers**



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## **Table of Contents Limiting Reagent Lab Answer Key**

1. Understanding the eBook Limiting Reagent Lab Answer Key
  - The Rise of Digital Reading Limiting Reagent Lab Answer Key
  - Advantages of eBooks Over Traditional Books
2. Identifying Limiting Reagent Lab Answer Key
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Limiting Reagent Lab Answer Key
  - User-Friendly Interface
4. Exploring eBook Recommendations from Limiting Reagent Lab Answer Key
  - Personalized Recommendations
  - Limiting Reagent Lab Answer Key User Reviews and Ratings
  - Limiting Reagent Lab Answer Key and Bestseller Lists
5. Accessing Limiting Reagent Lab Answer Key Free and Paid eBooks
  - Limiting Reagent Lab Answer Key Public Domain eBooks
  - Limiting Reagent Lab Answer Key eBook Subscription Services
  - Limiting Reagent Lab Answer Key Budget-Friendly Options

6. Navigating Limiting Reagent Lab Answer Key eBook Formats
  - ePub, PDF, MOBI, and More
  - Limiting Reagent Lab Answer Key Compatibility with Devices
  - Limiting Reagent Lab Answer Key Enhanced eBook Features
7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Limiting Reagent Lab Answer Key
  - Highlighting and Note-Taking Limiting Reagent Lab Answer Key
  - Interactive Elements Limiting Reagent Lab Answer Key
8. Staying Engaged with Limiting Reagent Lab Answer Key
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Limiting Reagent Lab Answer Key
9. Balancing eBooks and Physical Books Limiting Reagent Lab Answer Key
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Limiting Reagent Lab Answer Key
10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
11. Cultivating a Reading Routine Limiting Reagent Lab Answer Key
  - Setting Reading Goals Limiting Reagent Lab Answer Key
  - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Limiting Reagent Lab Answer Key
  - Fact-Checking eBook Content of Limiting Reagent Lab Answer Key
  - Distinguishing Credible Sources
13. Promoting Lifelong Learning
  - Utilizing eBooks for Skill Development
  - Exploring Educational eBooks
14. Embracing eBook Trends
  - Integration of Multimedia Elements

- Interactive and Gamified eBooks

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