

# Final study guide question and answers

What kind of equilibrium exists and under what conditions? **Chemical equilibrium** is a dynamic equilibrium between two or more states of matter, where the concentrations of the reactants and products are constant over time.

What happens if you change the conditions? **Le Chatelier's principle** states that if a system at equilibrium is subjected to a change in concentration, temperature, or pressure, the system will shift its equilibrium position to counteract the change.

What happens if you change the volume of the container? **Changing the volume** of a container at constant temperature and pressure will shift the equilibrium position towards the side with more moles of gas.

What happens if you change the temperature? **Changing the temperature** of a system at constant pressure and volume will shift the equilibrium position towards the endothermic direction.

What happens if you change the concentration of the reactants or products? **Changing the concentration** of a reactant or product will shift the equilibrium position towards the side with the lower concentration.

What happens if you change the pressure? **Changing the pressure** of a system at constant temperature and volume will shift the equilibrium position towards the side with fewer moles of gas.

What happens if you change the catalyst? **A catalyst** speeds up the rate at which equilibrium is reached but does not change the equilibrium position.

What happens if you change the initial concentrations? **Changing the initial concentrations** of a system at constant temperature and pressure will shift the equilibrium position towards the side with the higher initial concentration.

What happens if you change the initial volume? **Changing the initial volume** of a system at constant temperature and pressure will shift the equilibrium position towards the side with more moles of gas.

What happens if you change the initial temperature? **Changing the initial temperature** of a system at constant pressure and volume will shift the equilibrium position towards the endothermic direction.

What happens if you change the initial concentration of the reactants or products? **Changing the initial concentration** of a reactant or product will shift the equilibrium position towards the side with the lower initial concentration.

What happens if you change the initial volume of the container? **Changing the initial volume** of a container at constant temperature and pressure will shift the equilibrium position towards the side with more moles of gas.

What happens if you change the initial pressure? **Changing the initial pressure** of a system at constant temperature and volume will shift the equilibrium position towards the side with fewer moles of gas.

What happens if you change the initial catalyst? **A catalyst** speeds up the rate at which equilibrium is reached but does not change the equilibrium position.

What are some examples of chemical equilibrium? **Examples of chemical equilibrium** include the Haber process, the decomposition of calcium carbonate, and the reaction between nitrogen and hydrogen to form ammonia.

# Nate Quick Study Guide

**Michael Brown**



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