

Thermodynamics

MULTIPLE CHOICE On the line at the left of each statement, write the letter of the choice that best completes the statement or answers the question.

- _____ If an ideal gas does positive work on its surroundings,
 - the temperature of the gas increases.
 - the volume of the gas increases.
 - the pressure of the gas increases.
 - the internal energy of the gas decreases.
- _____ Sunlight strikes an ice cube at its melting point and causes it to melt. What is this process?

a. adiabatic process	e. isobaric process
b. isothermal process	d. isovolumetric process
- _____ Which of the following processes for an ideal gas system has an unchanging internal energy and a heat intake that corresponds to the value of the work done by the system?

a. isothermal process	e. adiabatic process
b. isobaric process	d. isovolumetric process
- _____ How is conservation of internal energy expressed for an isovolumetric system?
 - $Q = W = \Delta U = 0$; $U_i = U_f$
 - $Q = 0$; so $\Delta U = -W$
 - $\Delta T = 0$; so $\Delta U = Q - W = 0$; so $Q = W$
 - $\Delta V = 0$; so $P\Delta V = 0$ and $W = 0$; so $\Delta U = Q$
- _____ A 4 mol ideal gas system undergoes an adiabatic process in which it expands and does 20 J of work on its environment. What is its change in internal energy?

a. -20 J	e. 0 J
b. -5 J	d. 20 J
- _____ According to the second law of thermodynamics, the heat received by a heat engine operating in a complete cycle from a high-temperature reservoir
 - must be completely converted to work.
 - equals the entropy increase.
 - can be completely converted to internal energy.
 - cannot be completely converted to work.

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Gerard Gouesbet



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