

## Form 4 Chemistry Notes

Group 1	Group 15	$M_3N$	$Na_3N$
Group 1	Group 16	$M_2N$	$K_2O$
Group 1	Group 17	$MN$	$LiCl$
Group 2	Group 15	$M_3N_2$	$Ca_3N_2$
Group 2	Group 16	$MN$	$MgO$
Group 2	Group 17	$MN_2$	$BaBr_2$
Group 13	Group 15	$MN$	$AlN$
Group 13	Group 16	$M_2N_3$	$Al_2O_3$
Group 13	Group 17	$MN_3$	$AlI_3$

### Covalent Bonding

1. Covalent bonds are formed by atoms sharing electrons to form molecules. This type of bond usually formed between two non-metallic elements.
2. In the formation of covalent bonds, atoms of non-metals will combine with each other to donate one, two or three electrons for sharing.
3. The compound formed through the formation of covalent bonds is called the covalent compounds.

### Examples of Covalent Compound

Covalent Compound	Formula	Covalent Compound	Formula
Chlorine	$Cl_2$	Phosphorus	$P_4$
Hydrogen	$H_2$	Sulfur dioxide	$SO_2$
Oxygen	$O_2$	Carbon dioxide	$CO_2$

# Form Chemistry Notes

**John William Major Sothern**



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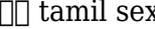
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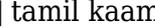
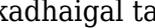
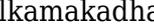
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