

## Chemistry - Solubility Curves Virtual Lab

**Part I:** Construct a single graph with multiple solubility curves to clearly show the solubility of the four ionic compounds in the data table below. The following tips will help you make the best possible graph.

1. Make the graph **LARGE**. It should fill at least three-fourths of the graph paper.
  2. The **Y-axis** should be labeled appropriately (with units) and have an appropriate range to accommodate the solubility of each solute (**Assume all the solutes are dissolved in 100. g H<sub>2</sub>O**)
  3. The **X-axis** should be labeled appropriately (with units) and a range of values to accommodate all the temperature values.
  4. The graph should have a descriptive title. (y-axis vs. x-axis)
  5. The solubility curve for each compound should be a different color, and make sure to write the compound's formula by the curve on the graph. (see bottom of Data table)
  6. Use the graph to extrapolate and/or interpolate numbers for the questions using the data table below. (**Do not fill any data in that is missing, just wait until you answer the questions!**)
1. What is the independent variable? [ **Temperature in celsius** ] What axis does this variable go on? [ **x-axis** ]
  2. What is the dependent variable? [ **The number of grams dissolved in 100g of H<sub>2</sub>O** ] What axis does this variable go on? [ **y-axis** ]

Solubility in grams of Solute dissolved in 100 grams of water vs. Temperature

Temperature (°C)	KNO <sub>3</sub> (grams)	KCl (grams)	Ba(OH) <sub>2</sub> (grams)	CO <sub>2</sub> (grams)
0.0°C	12 g		2 g	100. g
10. °C	19 g	30. g	3 g	80 g
20.°C	30. g		4 g	76 g
30.°C	42 g	36 g	6 g	47 g
40.°C	60. g		8 g	40. g
50.°C	80. g	42 g	12 g	37 g
60.°C	100. g		24 g	
70.°C	122 g	48 g	52 g	
80.°C	157 g		96 g	28 g
90.°C		55 g	145 g	
100.°C			215 g	
<b>Line Color</b>	<b>Green</b>	<b>Blue</b>	<b>Orange</b>	<b>Red</b>

**Part II:** Use the graph you made to answer the following questions. If the problem requires calculations, **SHOW ALL WORK** to receive credit for the answer. No work, no credit, no kidding. **Place the final answer to be graded in the blanks provided.**

1. a) List the formulas of the solutes above which are **SOLIDS**. [ **KCl, KNO<sub>3</sub> and Ba(OH)<sub>2</sub>** ]  
b) List the formulas of the solutes above which are **GASES**. [ **Co<sub>2</sub>** ]
2. a) What type of relationship exists between temperature and the solubility of CO<sub>2</sub>? [ **As the temperature increases there are less soluble.** ]  
b) Why does this relationship exist? [ **The higher the temperature the lower the solubility because the gas dissolves faster in lower temperatures.** ]

# Lab Solubility Data Sheet Answer Key

**LL Leslie**



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