

# ACID-BASE RELATIVE STRENGTHS in Aqueous Solutions at 25° C

Name of Acid	Acid		Base	$K_a$
Hydriodic	HI	→	H <sup>+</sup> + I <sup>-</sup>	≫ 1
Hydrobromic	HBr	→	H <sup>+</sup> + Br <sup>-</sup>	≫ 1
Hydrochloric	HCl	→	H <sup>+</sup> + Cl <sup>-</sup>	≫ 1
Nitric	HNO <sub>3</sub>	→	H <sup>+</sup> + NO <sub>3</sub> <sup>-</sup>	> 1
Sulfuric	H <sub>2</sub> SO <sub>4</sub>	→	H <sup>+</sup> + HSO <sub>4</sub> <sup>-</sup>	> 1
Oxalic	H <sub>2</sub> C <sub>2</sub> O <sub>4</sub>	⇌	H <sup>+</sup> + HC <sub>2</sub> O <sub>4</sub> <sup>-</sup>	5.90 × 10 <sup>-2</sup>
Sulfurous	H <sub>2</sub> SO <sub>3</sub>	⇌	H <sup>+</sup> + HSO <sub>3</sub> <sup>-</sup>	1.54 × 10 <sup>-2</sup>
Hydrogen Sulfate Ion	HSO <sub>4</sub> <sup>-</sup>	⇌	H <sup>+</sup> + SO <sub>4</sub> <sup>2-</sup>	1.20 × 10 <sup>-2</sup>
Phosphoric	H <sub>3</sub> PO <sub>4</sub>	⇌	H <sup>+</sup> + H <sub>2</sub> PO <sub>4</sub> <sup>-</sup>	7.52 × 10 <sup>-3</sup>
Nitrous	HNO <sub>2</sub>	⇌	H <sup>+</sup> + NO <sub>2</sub> <sup>-</sup>	4.6 × 10 <sup>-4</sup>
Hydrofluoric	HF	⇌	H <sup>+</sup> + F <sup>-</sup>	3.53 × 10 <sup>-4</sup>
Formic	HCO <sub>2</sub> H	⇌	H <sup>+</sup> + HCO <sub>2</sub> <sup>-</sup>	1.77 × 10 <sup>-4</sup>
Benzoic	C <sub>6</sub> H <sub>5</sub> CO <sub>2</sub> H	⇌	H <sup>+</sup> + C <sub>6</sub> H <sub>5</sub> CO <sub>2</sub> <sup>-</sup>	6.46 × 10 <sup>-5</sup>
Acetic	CH <sub>3</sub> CO <sub>2</sub> H	⇌	H <sup>+</sup> + CH <sub>3</sub> CO <sub>2</sub> <sup>-</sup>	1.76 × 10 <sup>-5</sup>
Carbonic	H <sub>2</sub> CO <sub>3</sub>	⇌	H <sup>+</sup> + HCO <sub>3</sub> <sup>-</sup>	4.30 × 10 <sup>-7</sup>
Hydrogen Sulfite Ion	HSO <sub>3</sub> <sup>-</sup>	⇌	H <sup>+</sup> + SO <sub>3</sub> <sup>2-</sup>	1.02 × 10 <sup>-7</sup>
Dihydrogen Phosphate Ion	H <sub>2</sub> PO <sub>4</sub> <sup>-</sup>	⇌	H <sup>+</sup> + HPO <sub>4</sub> <sup>2-</sup>	6.23 × 10 <sup>-8</sup>
Boric	H <sub>3</sub> BO <sub>3</sub>	⇌	H <sup>+</sup> + H <sub>2</sub> BO <sub>3</sub> <sup>-</sup>	7.3 × 10 <sup>-10</sup>
Ammonium Ion	NH <sub>4</sub> <sup>+</sup>	⇌	H <sup>+</sup> + NH <sub>4</sub> <sup>+</sup>	5.64 × 10 <sup>-10</sup>
Hydrogen Carbonate Ion	HCO <sub>3</sub> <sup>-</sup>	⇌	H <sup>+</sup> + CO <sub>3</sub> <sup>2-</sup>	5.61 × 10 <sup>-11</sup>
Monohydrogen Phosphate Ion	HPO <sub>4</sub> <sup>2-</sup>	⇌	H <sup>+</sup> + PO <sub>4</sub> <sup>3-</sup>	2.2 × 10 <sup>-13</sup>
Water	H <sub>2</sub> O	⇌	H <sup>+</sup> + OH <sup>-</sup>	1.0 × 10 <sup>-14</sup>

Strongest Acid

Weak Acid

Weakest Acid

Weakest Base

Weak Base

Strongest Base

For space considerations, all reactions omit water as a reactant  
 $HA + H_2O \rightleftharpoons H_3O^+ + A^-$

# Lab 16 1 Creating Acid Base Strength Table

**Brendan G. Carr**



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