

# ICE Practice Problems

#1 – Relatively easy, no ICE table required because eq'm concentrations are given

For the reaction  $\text{CH}_4(\text{g}) + \text{H}_2\text{O}(\text{g}) \leftrightarrow \text{CO}(\text{g}) + 3\text{H}_2(\text{g})$  @  $1500^\circ\text{C}$  an equilibrium mixture of these gases was found to have the following concentrations  $[\text{CO}] = 0.300\text{M}$ ,  $[\text{H}_2] = 0.800\text{M}$  and  $[\text{CH}_4] = 0.400\text{M}$ .  $K_c$  @  $1500^\circ\text{C} = 5.67$ . Determine the equilibrium concentration of  $\text{H}_2\text{O}$  in this mixture.

#2 – Requires an ICE table because you do not know the equilibrium concentrations- no product is yet formed

For the reaction  $\text{CO}(\text{g}) + \text{H}_2\text{O}(\text{g}) \leftrightarrow \text{CO}_2(\text{g}) + \text{H}_2(\text{g})$  calculate the equilibrium concentrations of all species if 1.000 mol of each reactant is mixed in a 1.000L flask.  $K_c = 5.10$  at the temperature of this reaction.

		$\text{CO}(\text{g})$	+	$\text{H}_2\text{O}(\text{g})$	$\leftrightarrow$	$\text{CO}_2(\text{g})$	+	$\text{H}_2(\text{g})$
[Initial]	I	1.000		1.000		0		0
[Change in]	C							
[Equilibrium]	E							

#3 – Requires an ICE table because you do not know the equilibrium concentrations. The initial concentrations must be calculated- no product is yet formed

For the reaction  $\text{H}_2(\text{g}) + \text{F}_2(\text{g}) \leftrightarrow 2\text{HF}(\text{g})$  calculate the equilibrium concentrations of all species if 3.000 mol of each reactant was added 1.500L flask.  $K_c$  at the temperature of the reaction is  $1.15 \times 10^2$ .

		$\text{H}_2(\text{g})$	+	$\text{F}_2(\text{g})$	$\leftrightarrow$	$2\text{HF}(\text{g})$
[Initial]	I	2.00		2.00		0
[Change in]	C					
[Equilibrium]	E					

#4 – Requires an ICE table because you do not know the equilibrium concentrations. Initial concentrations of reactants are given.

0.200mol of  $\text{H}_2$ , 0.200mol of  $\text{I}_2$ , and 0.200mol of  $\text{HI}$  were placed in a 1.00 L flask and allowed to come to equilibrium. The  $K_c$  value of the reaction at this temperature is 49.5. Determine the equilibrium concentrations of all species.

		$\text{H}_2(\text{g})$	+	$\text{I}_2(\text{g})$	$\leftrightarrow$	$2\text{HI}(\text{g})$
[Initial]	I					
[Change in]	C					
[Equilibrium]	E					

#5 – Requires an ICE table because you do not know the equilibrium concentrations. Initial concentrations of reactants must be calculated and no product is yet formed

For the reaction  $\text{H}_2(\text{g}) + \text{F}_2(\text{g}) \leftrightarrow 2\text{HF}(\text{g})$  calculate the equilibrium concentrations of each species if 3.000 mol of  $\text{H}_2$  and 6.000mol of  $\text{F}_2$  are mixed in a 3.000L flask.  $K_c$  at this temperature is  $1.15 \times 10^2$ .

		$\text{H}_2(\text{g})$	+	$\text{F}_2(\text{g})$	$\leftrightarrow$	$2\text{HF}(\text{g})$
[Initial]	I					
[Change in]	C					
[Equilibrium]	E					

# Ice Practice Problems Answer Key

**G Thomas**



## Ice Practice Problems Answer Key:

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## **Ice Practice Problems Answer Key** Book Review: Unveiling the Magic of Language

In a digital era where connections and knowledge reign supreme, the enchanting power of language has become more apparent than ever. Its ability to stir emotions, provoke thought, and instigate transformation is truly remarkable. This extraordinary book, aptly titled "**Ice Practice Problems Answer Key**," compiled by a highly acclaimed author, immerses readers in a captivating exploration of the significance of language and its profound affect our existence. Throughout this critique, we will delve to the book is central themes, evaluate its unique writing style, and assess its overall influence on its readership.

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