

Basic Conversion Factors: 6.02×10^{23} molecules or atoms = 1 mole of that substance

Molarity (#M) = # moles to 1L

Molar mass = #g to 1 mole of that substance

1000ml = 1L

Name _____

Period _____

Mixed Dimensional Analysis Review

Show your work, label all units of measurement throughout, and honor the multiplication/division rule for significant figures.

1. Convert 59.99g of Fe to moles.
2. Convert .39 moles of $\text{Cu}(\text{OH})_2$ to grams.
3. Convert 2.77×10^{24} molecules of NH_3 to moles.
4. Convert 1.9 moles of HBr to molecules.
5. Convert 93.16g of Mg to atoms.
6. Convert 8.20×10^{22} molecules of ZnCl_2 to grams.

Mixed Dimensional Analysis Worksheet

JG Myers



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