

Molarity, Dilutions, and Preparing Solutions

A. Preparing a Stock Solution of Known Concentration

1. Mass of 100-mL volumetric flask, g 59.45
 2. Mass of 100-mL volumetric flask and CuSO_4 , g 64.08
 3. Mass of CuSO_4 , g 4.63
 4. Volume of stock solution, mL 100 mL
 5. Molarity of stock solution "S," M 0.0290 M
- (Show calculations.)
- $$\frac{4.63}{159.6} = \frac{0.0290}{0.100}$$

B. Preparing the Spectrophotometer for use

1. Percent transmission of DI H_2O , %T 100 %
2. Absorbance of DI H_2O , A 0

C. Measuring the Absorbance of the Stock Solution

1. Percent transmission of stock solution, %T 43.9
2. Absorbance of stock solution, A 0.357

E. Construction of a Calibration Graph

	"A"	"B"	"C"	"D"
Volume of stock solution, mL	5.00	10.00	15.00	20.00
Volume of prepared solution, mL	50.00	50.00	50.00	50.00
Molarity, M	0.0058	0.0116	0.0174	0.0230
Percent transmission, %T	105	82.3	79.5	52.6
Absorbance, A	.021	.085	.100	.280

Molarity Concentration Lab Report

Lauren Gardner



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