

## Molarity Worksheet

Name: \_\_\_\_\_

1. What does molarity mean?

$$\frac{\text{Number of moles of solute}}{\text{1 liter solution}}$$

2. What is the molarity of a solution that contains 4.53 moles of lithium nitrate in 2.85 liters of solution?

$$\frac{4.53 \text{ mol LiNO}_3}{2.85 \text{ L soln}} = 1.59 \text{ M LiNO}_3$$

3. What is the molarity of a solution that contains 0.00372 moles hydrochloric acid in  $2.39 \times 10^{-2}$  liters of solution?

$$\frac{0.00372 \text{ mol HCl}}{2.39 \times 10^{-2} \text{ L soln}} = 0.156 \text{ M HCl}$$

4. A flask contains 85.5 g  $\text{C}_{12}\text{H}_{22}\text{O}_{11}$  (sucrose) in 1.00 liters of solution. What is the molarity?

$$\frac{85.5 \text{ g sucrose} \times 1 \text{ mol sucrose}}{1.00 \text{ L soln} \quad 342.34 \text{ g sucrose}} = 0.250 \text{ M sucrose}$$

5. A beaker contains 214.2 grams osmium (III) fluoride in 0.0673 liters of solution. What is the molarity?

$$\frac{214.2 \text{ g OsF}_3 \times 1 \text{ mol OsF}_3}{0.0673 \text{ L soln} \quad 247.23 \text{ g OsF}_3} = 12.9 \text{ M OsF}_3$$

6. Calculate the molarity if a flask contains 1.54 moles potassium sulfate in 125 ml of solution.

$$\frac{1.54 \text{ mol K}_2\text{SO}_4}{0.125 \text{ L soln}} = 12.3 \text{ M K}_2\text{SO}_4$$

7. A beaker contains 36.45 grams ammonium chlorite in 2.36 liters of solution - calculate the molarity.

$$\frac{36.45 \text{ g NH}_4\text{ClO}_2 \times 1 \text{ mol NH}_4\text{ClO}_2}{2.36 \text{ L soln} \quad 85.50 \text{ g NH}_4\text{ClO}_2} = 0.181 \text{ M NH}_4\text{ClO}_2$$

8. What is the molarity of a solution that contains 14.92 grams magnesium oxalate in 3.65 ml of solution?

$$\frac{14.92 \text{ g MgC}_2\text{O}_4 \times 1 \text{ mol MgC}_2\text{O}_4}{0.00365 \text{ L soln} \quad 112.32 \text{ g MgC}_2\text{O}_4} = 36.4 \text{ mol MgC}_2\text{O}_4$$

# Molarity Pre Lab Worksheet

**Didier Musso**



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